

DEPARTMENT OF CHEMISTRY

B.Sc. Chemistry Syllabus under CBCS Semester Wise Syllabus & Model Question Papers

WITH EFFECT FROM 2019-20



Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited **A** Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Board Of Studies October- 2018

Dr. V. S. Krishna Govt. Degree & P.G. College (A), Visakhapatnam
B.Sc. Chemistry Syllabus Under CBCS
[w.e.f. 2019-20 (Revised in October 2018)]

Year	Semester	Paper	Title	Marks	Credits	
I	I	I	Inorganic and Organic Chemistry	100	03	
			Practical – I	50	02	
	II	II	Physical and General Chemistry	100	03	
			Practical – II	50	02	
II	III	III	Inorganic and organic Chemistry	100	03	
			Practical – III	50	02	
	IV	IV	Spectroscopy and Physical Chemistry	100	03	
			Practical – IV	50	02	
III	V	V	Inorganic ,Organic and Physical Chemistry	100	03	
			Practical – V	50	02	
		VI	Inorganic ,Organic and Physical Chemistry	100	03	
			Practical – VI	50	02	
		* Any one Paper from VII A, B and C	VII (A)*	Elective	100	03
				Practical - VII A	50	02
	VII (B)*		Elective	100	03	
			Practical - VII B	50	02	
	VII (C)*		Elective	100	03	
			Practical - VII C	50	02	
	** Any one cluster from VIII, A, B and C	VIII (A)**	Cluster Electives - I :	100	03	
			VIII-A-1	100	03	
			VIII-A-2	100	03	
			VIII-A-3	50	02	
				50	02	
				50	02	
		VIII (B)**	Cluster Electives - II ::	100	03	
			VIII-B-1	100	03	
VIII- B-2			100	03		
VIII-B-3			50	02		
			50	02		
			50	02		
VI	VIII (C)**	Cluster Electives - III ::	100	03		
		VIII-C-1	100	03		
		VIII-C-2	100	03		
		VIII-C-3	50	02		
			50	02		
			50	02		

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CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- I, Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

SEMESTER – I

Paper I - Inorganic & Organic Chemistry 60hrs (4h/w)

INORGANIC CHEMISTRY

30 hrs (2h / w)

UNIT –I

p-block elements –I

15h

Group-13: Synthesis and structure of diborane and higher boranes (B_4H_{10} and B_5H_9), boron-nitrogen compounds ($B_3N_3H_6$ and BN)

Group - 14: Preparation and applications of silanes and silicones.

Group - 15: Preparation and reactions of hydrazine, hydroxylamine.

UNIT-II

1. p-block elements -II

8h

Group-16: Classifications of oxides based on (i) Chemical behavior and (ii) Oxygen content.

Group-17: Inter halogen compounds and pseudo halogens.

2. Organometallic Chemistry

7h

Definition- classification of Organometallic compounds- nomenclature, preparation, properties and applications of alkyls of Li and Mg.

ORGANIC CHEMISTRY

30hrs (2h /w)

UNIT-III

Structural theory in Organic Chemistry

10 h

Types of bond fission and organic reagents (Electrophilic, Nucleophilic, and free radical reagents including neutral molecules like H_2O , NH_3 & $AlCl_3$).

Bond polarization: Factors influencing the polarization of covalent bonds, electro negativity - inductive effect. Application of inductive effect (a) Basicity of amines (b) Acidity of carboxylic acids (c) Stability of carbonium ions. Resonance or Mesomeric effect, application to (a) acidity of phenol, and (b) acidity of carboxylic acids. Hyper conjugation and its application to stability of carbonium ions, Free radicals and alkenes, carbanions, carbenes and nitrenes.

Types of Organic reactions: Addition - electrophilic, nucleophilic and free radical. Substitution - electrophilic, nucleophilic and free radical. Elimination- Examples.

UNIT-IV

1. Acyclic Hydrocarbons

6 h

Alkenes - Preparation of alkenes. Properties: Addition of hydrogen - heat of hydrogenation and stability of alkenes. Addition of halogen and its mechanism. Addition of HX, Markonikov's rule, addition of H₂O, HOX, H₂SO₄ with mechanism and addition of HBr in the presence of peroxide (anti - Markonikov's addition). Dienes - Types of dienes, reactions of conjugated dienes - 1, 2 and 1, 4 addition of HBr to 1, 3 - butadiene and Diel's - Alder reaction.

Alkynes - Preparation by dehydrohalogenation of dihalides, dehalogenation of tetra halides, Properties; Acidity of acetylenic hydrogen (formation of Metal acetylides). Preparation of higher acetylenes, Metal ammonia reductions, Physical properties. Chemical reactivity-electrophilic addition of X₂, HX, H₂O (Tautomerism), Oxidation with KMnO₄, OsO₄, reduction and Polymerization reaction of acetylene.

2. Alicyclic hydrocarbons (Cycloalkanes)

4 h

Nomenclature, Preparation by Freund's method, Wislicenus method. Properties - reactivity of cyclopropane and cyclobutane by comparing with alkanes, Stability of cycloalkanes - Baeyer's strain theory, Sachse and Mohr predictions and Pitzer's strain theory. Conformational structures of cyclobutane, cyclopentane, cyclohexane.

UNIT-V

Benzene and its reactivity

10h

Concept of resonance, resonance energy. Heat of hydrogenation, heat of combustion of Benzene, mention of C-C bond lengths and orbital picture of Benzene. Concept of aromaticity - aromaticity (definition), Huckel's rule - application to Benzenoid (Benzene, Naphthalene) and Non - Benzenoid compounds (cyclopropenyl cation, cyclopentadienyl anion and tropylium cation)

Reactions - General mechanism of electrophilic substitution, mechanism of nitration, Friedel Craft's alkylation and acylation. Orientation of aromatic substitution - Definition of ortho, para and meta directing groups. Ring activating and deactivating groups with examples (Electronic interpretation of various groups like NO₂ and Phenolic). Orientation of (i) Amino, methoxy and methyl groups (ii) Carboxy, nitro, nitrile, carbonyl and sulphonic acid groups (iii) Halogens
(Explanation by taking minimum of one example from each type)

List of Reference Books

1. Inorganic Chemistry by J.E.Huheey
2. Basic Inorganic Chemistry by Cotton and Wilkinson
3. A textbook of qualitative inorganic analysis by A.I. Vogel
4. Organic Chemistry by Morrison and Boyd
5. A Text Book of Organic chemistry by I L Finar Vol I
6. Concise Inorganic Chemistry by J.D.Lee

Guidelines to the Paper Setter: The syllabus I Semester consists of **Inorganic & Organic chemistry**. The I Semester question paper consists of 2 sections.

In Section- A: Consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from Inorganic & organic part

In Section- B: Consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 3 questions from Inorganic & 5 questions from organic part.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

B. Sc Chemistry, Paper-I, Semester-I
BLUE PRINT
(Inorganic & Organic Chemistry)

Sl. No	UNITS	Name of the chapter	8 Marks	4 Marks
INORGANIC CHEMISTRY				
1	UNIT-I	p-block elements –I	1	1
2	UNIT-II	p-block elements -II	2	1
		Organometallic Chemistry	1	1
ORGANIC CHEMISTRY				
3	UNIT-III	Structural theory in Organic Chemistry	2	1
4	UNIT-IV	Acyclic Hydrocarbons	1	1
		Alicyclic hydrocarbons (Cycloalkanes)	1	1
5	UNIT-V	Benzene and its reactivity	2	2

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)

B. Sc, DEGREE FIRST YEAR EXAMINATIONS

Paper –I, SEMESTER- I

(INORGANIC & ORGANIC CHEMISTRY)

Time: 3 hours

Maximum Marks: 60Marks

PART- A

Answer ALL the questions

5x8 = 40 Marks

1. (a) Explain the Preparation, Structure and Properties of Silicones.

సిలికొన్లు తయారీ, నిర్మాణం మరియు ధర్మాలు ను వివరించండి.

(OR)

- (b) Define the electron deficient molecules and draw the structure of Diborane.

ఎలక్ట్రాన్ కొరత అణువులు అనాగనేమి? డైబోరేన్ నిర్మాణాన్ని వ్రాయండి?

2. (a) Explain the different types of interhalogen compounds.

అంతర హాలోజన్ సమ్మేళనాలు రకాలును వివరించండి.

(OR)

- (b) Define and classify the organometallic compounds, and write their synthetic applications.

కర్బన లోహ సమ్మేళనాలు అనగా నిర్వచించి, వాటి వర్గీకరణను వివరించి సంశ్లేషిత అనువర్తనాలను తెల్పండి

3. (a) Describe different types of Organic Reactions with one example.

వివద కర్బన రసాయన చర్యలును ను ఉదాహరణతో వివరించండి.

(OR)

- (b) Write notes on the following. ఈ క్రింది వాటిని వివరించండి.

1) Mesomeric effect - మీసోమెరిక్ ప్రభావం 2) Hyper conjugation - అతిసంయుగ్మ

ప్రభావం 3) Inductive effect - ప్రేరేపక ప్రభావం

4. (a) Explain the preparation and properties of alkenes.

ఆల్కీన్లు తయారీ మరియు ధర్మాలును వివరించండి.

(OR)

- (b) Explain Baeyer's bond angle strain theory.

బేయర్ ప్రయాస కోణ సిద్ధాంతం ను వివరించండి.

5. (a) Describe the structure and properties of Benzene.

బెంజీన్ నిర్మాణాన్ని మరియు ధర్మాలు వివరించండి.

(OR)

(b) Explain electrophilic substitution mechanism in Benzene in respect of nitration and Friedel craft's acylation?

బెంజీన్ లో జరిగే ఎలక్ట్రోఫిలిక్ ప్రతిక్షేపన చర్య విధానాలను నైట్రిషన్ మరియు ఫ్రీడెల్ క్రాఫ్ట్ యస్ట్రోషన్ లకు వివరించండి.

PART- B

Answer any **FIVE** of the following questions

5x4 = 20 Marks

6. Explain the Preparation and reactions of Hydroxylamine.
హైడ్రాక్సిల్ అమీన్ తయారీ, చర్యలును వివరించండి.
7. Classify the Oxides based on the oxygen content with one example.
ఆక్సిజన్ పరిమాణం ఆధారంగా ఆక్సైడ్లను ఒక ఉదాహరణతో వర్గీకరించండి.
8. Write any two synthesis applications of Organo Lithium Compounds.
ఆర్గానో లిథియం సమ్మేళనాలు యొక్క రెండు సంశ్లేషణ అనువర్తనములు వ్రాయండి.
9. Explain the Bond fission in organic Reactions.
కర్బన రసాయన శాస్త్రము లో బంధ విచ్ఛిన్నం వివరించండి.
10. Explain the Markonikov's rule with Example.
మార్కోవ్వికోవ్ నియమం ను ఉదాహరణతో వివరించండి
11. Explain the acidity of the Acetylinic hydrogen with example.
ఎసిటిలిన్ హైడ్రోజన్ యొక్క అమ్లత్వ ధర్మాన్ని ఉదాహరణతో వివరించండి.
12. Explain Friedel Crafts Alkylation reaction with Mechanism.
ఫ్రీడెల్ క్రాఫ్ట్ అల్కైలేషన్ చర్యను చర్య క్రమ విధానంతో వివరించండి.
13. Define aromaticity and apply the Huckel's rule to benzene and naphthalene.
ఎరోమాటిసిటీ ని నిర్వచించి మరియు హుకెల్ నియమాన్ని బెంజీన్ మరియు నాఫ్థాలీన్ కు వర్తింపజేయండి.

LABORATORY COURSE-I

Practical-I Simple Salt Analysis

30 hrs (2 h / w)

(At the end of Semester-I)

Qualitative inorganic analysis

Analysis of simple salt containing one anion and cation from the following

Anions: Carbonate, sulphate, chloride, bromide, acetate, nitrate, borate, phosphate.

Cations: Lead, copper, iron, aluminum, zinc, manganese, nickel, calcium, strontium, barium, potassium and ammonium.

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CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- II, Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

Paper II (Physical & General Chemistry)

60 hrs. (4h/w)

PHYSICAL CHEMISTRY

30 hrs (2h / w)

UNIT-I

Solid state

10h

Symmetry in crystals. Law of constancy of interfacial angles. The law of rationality of indices. The law of symmetry. Definition of lattice point, space lattice, unit cell. Bravais lattices and crystal systems. X-ray diffraction and crystal structure. Bragg's law. Defects in crystals. Stoichiometric and non-stoichiometric defects.

UNIT-II

1. Gaseous state

6 h

Compression factors, deviation of real gases from ideal behavior. Vander Waal's equation of state. P-V Isotherms of real gases, Andrew's isotherms of carbon dioxide, continuity of state. Critical phenomena. The vander Waal's equation and the critical state. Law of corresponding states. Relationship between critical constants and vander Waal's constants. Joule Thomson effect.

2. Liquid state

4 h

Structural differences between solids, liquids and gases. Liquid crystals, the mesomorphic state. Classification of liquid crystals into Smectic and Nematic. Differences between liquid crystal and solid/liquid. Application of liquid crystals as LCD devices.

UNIT-III

Solutions

10h

Liquid-liquid - ideal solutions, Raoult's law. Ideally dilute solutions, Henry's law. Non-ideal solutions. Vapour pressure - composition and vapour pressure- temperature curves. Azeotropes-HCl-H₂O, ethanol-water systems and fractional distillation. Partially miscible liquids-phenol-water, trimethylamine-water, nicotine-water systems. Effect of impurity on consolute temperature. Immiscible liquids and steam distillation.

Nernst distribution law. Calculation of the partition coefficient. Applications of distribution law.

UNIT-IV**1. Surface chemistry****8 h**

Definition of colloids. Solids in liquids(sols), preparation, purification, properties - kinetic, optical, electrical. Stability of colloids, Hardy-Schulze law, protective colloid.

Liquids in liquids (emulsions) preparation, properties, uses. Liquids in solids (gels) preparation, uses.

Adsorption: Physical adsorption, chemisorption. Freundlich, Langmuir adsorption isotherms. Applications of adsorption

2. Chemical Bonding**7h**

Valence bond theory, hybridization, VB theory as applied to ClF_3 , $\text{Ni}(\text{CO})_4$, Molecular orbital theory - LCAO method, construction of M.O. diagrams for homo-nuclear and hetero-nuclear diatomic molecules (N_2 , O_2 , CO and NO).

UNIT-V**Stereochemistry of carbon compounds****15 h**

Molecular representations- Wedge, Fischer, Newman and Saw-Horse formulae.

Optical isomerism: Optical activity- wave nature of light, plane polarized light, optical rotation and specific rotation.

Chiral molecules- definition and criteria (Symmetry elements) - Definition of enantiomers and diastereomers – Explanation of optical isomerism with examples Glyceraldehydes, Lactic acid, Alanine, Tartaric acid, 2, 3-dibromopentane.

D, L and R, S configuration methods and E, Z- configuration with examples.

List of Reference Books

1. Principles of physical chemistry by Prutton and Marron
2. Solid State Chemistry and its applications by Anthony R. West
3. Text book of physical chemistry by K L Kapoor
4. Text book of physical chemistry by S Glasstone
5. Stereochemistry of Organic compounds by E L Eliel
6. Advanced Organic Chemistry by F A Carey and R J Sundberg
7. Stereochemistry by P.S.Kalsi
8. Stereochemistry of Organic compounds by D. Nasipuri
9. Advanced physical chemistry by Bahl and Tuli
10. Advanced Inorganic Chemistry Vol-I by Satyaprakash, Tuli, Basu and Madan

Guidelines to the Paper Setter: The syllabus II Semester consists of **Physical & General Chemistry**. The II Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from Physical & General Chemistry part.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 4 questions from Physical & 4 questions from General Chemistry part.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

**B. Sc Chemistry, Paper-I, Semester-II
BLUE PRINT
(Physical & General Chemistry)**

Sl. No	Units	Name of the chapter	8 Marks	4 Marks
PHYSICAL CHEMISTRY				
1	UNIT-I	Solid state	2	1
2	UNIT-II	Gaseous state	2	1
		Liquid state	-	1
3	UNIT-III	Solutions	2	1
GENERAL CHEMISTRY				
4	UNIT-IV	Surface chemistry	1	1
		Chemical Bonding	1	1
5	UNIT-V	Stereochemistry of carbon compounds	2	2

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)

B. Sc, DEGREE FIRST YEAR EXAMINATIONS

Paper –I, SEMESTER– II

(PHYSICAL & GENERAL CHEMISTRY)

(w.e.f 2019-20 revised on October-2018)

Time: 3 hours

Maximum Marks: 60Marks

PART- A

Answer ALL the questions

5 x 8 = 40 Marks

1. (a) Define semi conductors? Write a detailed note on n-type and p-type semi conductors.
అర్ధ వాహకాలు ను నిర్వచించండి? n-రకం మరియు p-రకం అర్ధ వాహకాలు గురించి వివరించండి.

Or

Derive the Brags equation of crystals.
స్పటికాలు యొక్క బ్రాగ్ సమీకరణంను రాబట్టండి.

2. Derive the Vanderwall's equation of state for real gases? How does it explain the behavior of real gases?
నిజ వాయువులకు వాండర్వాల్స్ సమీకరణంను రాబట్టండి. ఇది నిజ వాయువుల యొక్క స్వభావాన్ని ఏ విధముగా వివరించినది.

Or

Derive the relationship between Vanderwaals constants and Critical Constants.
వాండర్వాల్స్ స్థిరాంకాలు మరియు సందిగ్ధ స్థిరాంకాలు మధ్య సంబంధాన్ని రాబట్టండి.

3. Explain Nernst distribution law and its applications.
నెర్న్ స్ట్ వితరణ నియమము ను మరియు వాటి అనువర్తనములను వివరించండి.

Or

What is critical solution temperature? Explain the phenol-water system in terms of C.S.T.?

సందిగ్ధ ద్రావణ ఉష్ణోగ్రత అనగా నేమి? ఫినోల్- నీరు యొక్క C.S.T ని వివరించండి.

4. (a) Explain the paramagnetic behavior of oxygen with the help of molecular orbital energy-diagram.
అణు ఆర్బిటాల్ శక్తి పటం ఆధారంగా ఆక్సిజన్ పారా అయిస్కాంత ధర్మాన్ని వివరించండి.

Or

(a) Write any two methods of preparations of colloids?
కొల్లాయిడులు యొక్క ఏవైనా రెండు తయారీలను వ్రాయండి.

(b) Explain Kinetic and Optical properties of colloids?

కొల్లాయడులు యొక్క గతిక మరియు కాంతి దర్మాలును వివరించండి.

5. Explain the CIP Rules to assign the R, S configuration with examples.
R, S విన్యాసాన్ని నిర్ణయించే CIP నియమాలు ను వివరించండి.

Or

Explain enantiomerism and diastereomerism with suitable examples.
ఎనన్సియోమెరిజం మరియు డయాస్టీరీయోమెరిజంలను ఉదాహరణలతో వివరించండి.

PART- B

Answer any **FIVE** of the following questions

5 x 4 = 20 Marks

6. Write a short note on schottky and frenkel defects in crystals?
స్పటికాలలో షాట్కీ మరియు ఫ్రెనికల్ లోపాలు ను వివరించండి.
7. What are liquid crystals? Write about their classification?
ద్రవ స్పటికాలు అనగా నేమి. వాటి యొక్క వర్గీకరణను వ్రాయండి.
8. Write a short note on Joule Thomson Effect?
జౌల్ థామ్సన్ ప్రభావం గురించి వ్రాయండి.
9. State and explain Raoult's Law.
రౌల్ట్ నియమం ను నిర్వచించి, వివరించండి.
10. What are protective colloids? Explain and give their uses?
రక్షక కొల్లాయడులు అనగానేమి? వాటి ఉపయోగాలు ని వివరించండి.
11. Discuss the structure of with the help of valance bond theory.
 $Ni(CO)_4$ నిర్మాణాన్ని వెలన్సి బంధ సిద్ధాంత ఆధారముగా వివరించండి.
12. Write a short note on molecular representation in organic molecules.
కర్బన అణువులలో అణువుల ప్రాతినిధ్యం పై చిన్న వ్యాసం రాయండి.
13. Explain E, Z configuration with Examples.
E, Z విన్యాసాన్ని ఉదాహరణలతో వివరించండి.

Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited A Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- III, Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

Paper III (INORGANIC & ORGANIC CHEMISTRY) 60 hrs (4 h / w)

INORGANIC CHEMISTRY 30 hrs (2h / w)

UNIT –I

1. Chemistry of d-block elements: 9h

Characteristics of d-block elements with special reference to electronic configuration, variable valence, magnetic properties, catalytic properties and ability to form complexes. Stability of various oxidation states

2. Theories of bonding in metals:

6h

Metallic properties and its limitations, Valence bond theory, Free electron theory, Explanation of thermal and electrical conductivity of metals, limitations, Band theory, formation of bands, explanation of conductors, semiconductors and insulators.

UNIT – II

3. Metal carbonyls : 7h

EAN rule, classification of metal carbonyls, structures and shapes of metal carbonyls of V, Cr, Mn, Fe, Co and Ni.

4. Chemistry of f-block elements: 8h

Chemistry of lanthanides - electronic structure, oxidation states, lanthanide contraction, consequences of lanthanide contraction, magnetic properties. Chemistry of actinides - electronic configuration, oxidation states, actinide contraction, comparison of lanthanides and actinides.

ORGANIC CHEMISTRY 30 h (2h/w)

UNIT – III

1. Halogen compounds 5 h

Nomenclature and classification of alkyl (into primary, secondary, tertiary), aryl, aryl alkyl, allyl, vinyl, benzyl halides.

Nucleophilic aliphatic substitution reaction- classification into S_N^1 and S_N^2 – reaction mechanism with examples – Ethyl chloride, t-butyl chloride and optically active alkyl halide 2-bromobutane.

2. Hydroxy compounds

5 h

Nomenclature and classification of hydroxy compounds.

Alcohols: Preparation with hydroboration reaction, Grignard synthesis of alcohols.

Phenols: Preparation i) from diazonium salt, ii) from aryl sulphonates, iii) from cumene.

Physical properties- Hydrogen bonding (intermolecular and intramolecular). Effect of hydrogen bonding on boiling point and solubility in water.

Identification of alcohols by oxidation with $KMnO_4$, Ceric ammonium nitrate, Luca's reagent and phenols by reaction with $FeCl_3$.

Chemical properties:

- Dehydration of alcohols.
- Oxidation of alcohols by CrO_3 , $KMnO_4$.
- Special reaction of phenols: Bromination, Kolbe-Schmidt reaction, Riemer-Tiemann reaction, Fries rearrangement, azocoupling, Pinacol-Pinacolone rearrangement.

UNIT-IV

Carbonyl compounds

10 h

Nomenclature of aliphatic and aromatic carbonyl compounds, structure of the carbonyl group. Synthesis of aldehydes from acid chlorides, synthesis of aldehydes and ketones using 1,3-dithianes, synthesis of ketones from nitriles and from carboxylic acids. Physical properties: Reactivity of carbonyl group in aldehydes and ketones.

Nucleophilic addition reaction with a) $NaHSO_3$, b) HCN , c) $RMgX$, d) NH_2OH , e) $PhNHNH_2$, f) 2,4 DNPH, g) Alcohols-formation of hemiacetal and acetal. Base catalysed reactions: a) Aldol, b) Cannizzaro's reaction, c) Perkin reaction, d) Benzoin condensation, e) Haloform reaction, f) Knoevenagel reaction. Oxidation of aldehydes- Baeyer-Villiger oxidation of ketones. Reduction: Clemmensen reduction, Wolf-Kishner reduction, MPV reduction, reduction with $LiAlH_4$ and $NaBH_4$. Analysis of aldehydes and ketones with a) 2,4-DNPH test, b) Tollen's test, c) Fehling test, d) Schiff's test e) Haloform test (with equation)

UNIT-V

1. Carboxylic acids and derivatives

6 h

Nomenclature, classification and structure of carboxylic acids. Methods of preparation by a) Hydrolysis of nitriles, amides b) Hydrolysis of esters by acids and bases with mechanism c) Carbonation of Grignard reagents. Special methods of preparation of aromatic acids by a) Oxidation of side chain. b) Hydrolysis by benzotrichlorides. c) Kolbe reaction. **Physical properties:** Hydrogen bonding, dimeric association, acidity-strength of acids with examples of trimethyl acetic acid and trichloroacetic acid. Relative differences in the acidities of aromatic and aliphatic acids. **Chemical properties:** Reactions involving H, OH and COOH groups- salt formation, anhydride formation, acid

chloride formation, amide formation and esterification (mechanism). Degradation of carboxylic acids by Huns-Diecker reaction, decarboxylation by Schimdt reaction, Arndt-Eistert synthesis, halogenation by Hell- Volhard- Zelinsky reaction.

2. Active methylene compounds

4 h

Acetoacetic ester: keto-enol tautomerism, preparation by Claisen condensation, Acid hydrolysis and ketonic hydrolysis. Preparation of a) monocarboxylic acids. b) Dicarboxylic acids. c) Reaction with urea

Malonic ester: preparation from acetic acid. **Synthetic applications:** Preparation of a) monocarboxylic acids (propionic acid and n-butyric acid). b) Dicarboxylic acids (succinic acid and adipic acid) c) α,β -unsaturated carboxylic acids (crotonic acid). d) Reaction with urea.

List of Reference Books

1. Selected topics in inorganic chemistry by W.D.Malik, G..D.Tuli,R.D.Madan
2. Inorganic Chemistry J E Huheey, E A Keiter and R L Keiter
3. A Text Book of Organic Chemistry by Bahl and Arun bahl
4. A Text Book of Organic chemistry by I L Finar Vol I
5. Organic chemistry by Bruice
6. Organic chemistry by Clayden
7. Advanced Inorganic chemistry by Gurudeep Raj
8. Basic Inorganic Chemistry by Cotton and Wilkinson
9. Concise Inorganic Chemistry by J.D.Lee

LABORATORY COURSE -III

30 hrs. (2 h / w)

Practical Paper-III Titrimetric analysis and Organic Functional Group Reactions (At the end of Semester-III)

Titrimetric analysis:

25M

1. Determination of Fe (II) using KMnO_4 with oxalic acid as primary standard.
2. Determination of Cu(II) using $\text{Na}_2\text{S}_2\text{O}_3$ with $\text{K}_2\text{Cr}_2\text{O}_7$ as primary standard.

Organic Functional Group Reactions

25M

3. Reactions of the following functional groups present in organic compounds (at least four) Alcohols, Phenols, Aldehydes, Ketones, Carboxylic acids and Amides

Guidelines to the Paper Setter: The syllabus III Semester consists of Inorganic & Organic chemistry. The III Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from Inorganic & organic part

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 3 questions from Inorganic & 5 questions from organic part.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

II B.Sc, Chemistry, III Semester				
BLUE PRINT				
(Inorganic & Organic chemistry)				
Sl. No	Units	Name of the chapter	8 Marks	4 Marks
INORGANIC CHEMISTRY				
1	UNIT-I	Chemistry of d-block elements	1	1
		Theories of bonding in metals	1	1
2	UNIT-II	Metal carbonyls	1	-
		Chemistry of f-block elements	1	1
ORGANIC CHEMISTRY				
3	UNIT-III	Halogen compounds	1	1
		Hydroxyl compounds	1	1
4	UNIT-IV	Carbonyl compounds	2	1
5	UNIT-V	Carboxylic acids and derivatives	1	1
		Active methylene compounds	1	1

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)
B. Sc, DEGREE SECOND YEAR EXAMINATIONS
Paper –III, SEMESTER– III
(INORGANIC & ORGANIC CHEMISTRY)
(w.e.f 2019-20 revised on October-2018)

Time: 3 hours

Maximum Marks: 60Marks

PART- A

Answer all FIVE questions. Each question carries 8 marks. 5 X 8 = 40 Marks
అన్ని ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు ఎనిమిది మార్కులు

1. What are transition elements? What are transition elements? Describe their variable oxidation states, Magnetic
పరివర్తన మూలకాల అనగా నేమి ? d-బ్లాక్ మూలకాల ఆక్సీకరణ స్థితులు, అయస్కాంత ధర్మాలు మరియు సంశ్లేష్ట సమ్మేళనాలను ఏర్పరిచే ప్రవృత్తులను వివరించుము.

OR

Explain Band theory of metals?
బంధసిద్ధాంతమును వివరించుము.

2. What are metal carbonyls? Explain the classification of metal carbonyls?
మెటల్ కార్బోనైల్స్ అనగా నేమి? మెటల్ కార్బోనైల్ ల వర్గీకరణను వివరించుము.

OR

What is lanthanide contraction? Explain the consequences of lanthanide Contraction?

లాంథ్ నైడు సంకోచం అనగానేమి ? ? లాంథ్ నైడుల సంకోచానికి గల కారణాలను వ్రాయండి..

3. Explain SN¹ and SN² reaction mechanism with the help of energy profile diagram.
SN¹ మరియు SN² చర్యలను, చర్యావిధానంను శక్తి పట సహాయంతో వివరించుము.

OR

Write a brief note on the following. క్రింది వాటిని వివరించుము.

- a) Acidic nature of phenols b) Riemer-Tiemann reaction c) Kolbe's reaction
a) ఫీనోల్ ల ఆమ్ల స్వభావం b) రైమర్-టీమన్ చర్య c) కోల్బ్ చర్య

4. Write any three methods of preparation of aldehydes.
ఏవైన మూడు ఆల్డిహైడ్ ల తయారీ విధానాలను వివరించుము.

OR

Explain the following reactions. . క్రింది చర్యలను వివరించుము

- a) Aldol reaction b) Cannizzaro's reaction c) Perkin reaction
a) ఆల్డోల్ చర్య b) కానిజారో చర్య c) పెర్కిన్ చర్య

5. Write any three methods of preparation of carboxylic acids.

ఏవైన మూడు కార్బాక్సిలిక్ ఆమ్లంల తయారీ విధానంల ను వివరించుము .

OR

How do you prepare Aceto acetic ester and write any two synthetic applications of it.

ఏసిటోఏసిటిక్ ఎస్టర్ తయారీని తెలిపి, ఏవైన రెండు సంశ్లేషణ అనువర్తనము లను వివరించండి .

SECTION – B

Answer any FIVE questions. Each question carries 4 marks 5 X 4 = 20 Marks

ఏవైనా ఐదు ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు నాలుగు మార్కులు

6. Explain catalytic properties of d-block elements.
d- బ్లాక్ మూలకాల ఉత్ప్రేరక ధర్మాలను వ్రాయండి
7. Write a short note on conductors and semi conductors.
వాహకాలు , అర్ధవాహకాలు గూర్చి వ్రాయండి.
8. Discuss structure of Nickel carbonyl and Iron carbonyl.
నికెల్ కార్బోనైల్ మరియు ఐరన్ కార్బోనైల్ నిర్మాణాలను వ్రాయండి
9. Explain nucleophilic substitution in 2-bromo butane.
2-బ్రోమో బ్యూటేన్ లో న్యుక్లియోఫిలిక్ ప్రతిక్షేపణను వివరించండి.
10. Write any two methods of Identification of alcohols.
ఏవైనా రెండు ఆల్కహాల్ ల గుర్తించుట కు పరీక్షలు వ్రాయండి.
11. Write any two reduction reactions of aldehydes.
ఏవైనా రెండు ఆల్డిహైడ్ ల ను గుర్తించుట కు పరీక్షలు వ్రాయండి.
12. Explain Huns-Dieckmann and HVZ reactions
హన్స్ డెకర్ మరియు HVZ చర్యలను వ్రాయండి.
13. Write any two applications of Malonic ester.
ఏవైనా రెండు మెలనిక్ ఎస్టర్ అనువర్తనాలను వ్రాయండి.

Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited A Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- IV, Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

Paper IV (SPECTROSCOPY & PHYSICAL CHEMISTRY) 60 hrs (4 h / w)

SPECTROSCOPY

30 hrs (2h / w)

UNIT-I

6h

General features of absorption - Beer-Lambert's law and its limitations, transmittance, Absorbance, and molar absorptivity. Single and double beam spectrophotometers. Application of Beer-Lambert law for quantitative analysis of 1. Chromium in $K_2Cr_2O_7$
2. Manganese in Manganous sulphate

Electronic spectroscopy:

8h

Interaction of electromagnetic radiation with molecules and types of molecular spectra. Energy levels of molecular orbitals (σ , π , n). Selection rules for electronic spectra. Types of electronic transitions in molecules effect of conjugation. Concept of chromophore and auxochrome.

UNIT-II

Infra red spectroscopy

8h

Different Regions in Infrared radiations. Modes of vibrations in diatomic and polyatomic molecules. Characteristic absorption bands of various functional groups. Interpretation of spectra-Alkanes, Aromatic, Alcohols carbonyls, and amines with one example to each.

Proton magnetic resonance spectroscopy (1H -NMR)

8h

Principles of nuclear magnetic resonance, equivalent and non-equivalent protons, position of signals. Chemical shift, NMR splitting of signals - spin-spin coupling, coupling constants. Applications of NMR with suitable examples-ethyl bromide, ethanol, acetaldehyde, 1, 1, 2-tribromo ethane, ethyl acetate, toluene and acetophenone.

PHYSICAL CHEMISTRY

30 hrs (2h / w)

UNIT-III

Dilute solutions

10h

Colligative properties. Raoult's law, relative lowering of vapour pressure, its relation to molecular weight of non-volatile solute. Elevation of boiling point and depression of freezing point. Derivation of relation between molecular weight and elevation in boiling point and depression in freezing point. Experimental methods of determination. Osmosis, osmotic pressure, experimental determination. Theory of dilute solutions. Determination

of molecular weight of non-volatile solute from osmotic pressure. Abnormal Colligative properties- Van't Hoff factor.

UNIT-IV

Electrochemistry-I

10h

Specific conductance, equivalent conductance. Variation of equivalent conductance with dilution. Migration of ions, Kohlrausch's law. Arrhenius theory of electrolyte dissociation and its limitations. Ostwald's dilution law. Debye-Huckel-Onsagar's equation for strong electrolytes (elementary treatment only). Definition of transport number, determination by Hittorfs method. Application of conductivity measurements- conductometric titrations.

UNIT-V

1. Electrochemistry-II

4h

Single electrode potential, sign convention, Reversible and irreversible cells Nernst Equation- Reference electrode, Standard Hydrogen electrode, calomel electrode, Indicator electrode, metal – metal ion electrode, Inert electrode, Determination of EMF of cell, Applications of EMF measurements - Potentiometric titrations.

2. Phase rule

6h

Concept of phase, components, degrees of freedom. Thermodynamic Derivation of Gibbs phase rule. Phase equilibrium of one component system - water system. Phase equilibrium of two- component system, solid-liquid equilibrium. Simple eutectic diagram of Pb-Ag system, simple eutectic diagram, desilverisation of lead, NaCl-Water system, Freezing mixtures.

List of Reference Books

1. Spectroscopy by William Kemp
2. Spectroscopy by Pavia
3. Organic Spectroscopy by J. R. Dyer
4. Modern Electrochemistry by J.O. M. Bockris and A.K.N.Reddy
5. Advanced Physical Chemistry by Atkins
6. Introduction to Electrochemistry by S. Glasstone
7. Elementary organic spectroscopy by Y.R. Sharma
8. Spectroscopy by P.S.Kalsi

Guidelines to the Paper Setter: The syllabus IV Semester consists of **Spectroscopy & Physical chemistry**. The IV Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from Spectroscopy & Physical chemistry part.

In Section - B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 4 questions from SPECTROSCOPY & 4 questions from PHYSICAL CHEMISTRY part.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

SEMESTER- IV II B.Sc, Chemistry Paper IV BLUE PRINT (Spectroscopy & Physical chemistry)				
Sl. No	Units	Name of the chapter	8 Marks	4 Marks
SPECTROSCOPY				
1	UNIT-I	General features of Spectroscopy	1	1
		Electronic spectroscopy	1	1
2	UNIT-II	Infra red spectroscopy	1	1
		Proton magnetic resonance spectroscopy	1	1
PHYSICAL CHEMISTRY				
3	UNIT-III	Dilute solutions	2	1
4	UNIT-IV	Electrochemistry-I	2	1
5	UNIT-V	Electrochemistry-II	1	1
		Phase rule	1	1

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)

B. Sc, DEGREE SECOND YEAR EXAMINATIONS

Paper –IV, SEMESTER– IV

(SPECTROSCOPY & PHYSICAL CHEMISTRY)

(w.e.f 2019-20 revised on October-2018)

Time: 3 hours

Maximum Marks: 60Marks

PART- A

Answer all FIVE questions. Each question carries 8 marks 5 x 8 = 40 Marks

అన్ని ప్రశ్నలకు సమాధానాలు వ్రాయండి..ప్రతి ప్రశ్నకు ఎనిమిది మార్కులు.

1. Explain single and double beam spectrophotometers.

ఏక పుంజ మరియు ద్వి పుంజ వర్ణ పట కాంతి మాపకాలను గూర్చి వ్రాయండి.

OR

Explain different types of electronic transitions in molecules with examples.

వివిధ రకాల అణువులో జరిగే ఎలక్ట్రాన్ పరివర్తనాలను ఉదాహరణలతో వివరించండి.

2. Explain different modes of vibrations in poly atomic molecules.

బహు పరమాణుక అణువులలో కంపన విధానాలను రెండు ఉదాహరణల తో వివరించుము.

OR

Explain the following.

క్రింది వాటిని వివరించుము.

a) Equivalent, nonequivalent protons b) Chemical shift c) Spin-Spin coupling

a) సమతుల్య, అసమతుల్య ప్రోటానులు b) రసాయన విస్థాపనం c) స్పిన్ - స్పిన్ విభజన

3. Explain the following.

క్రింది వాటిని వివరించుము.

a) Colligative properties with examples. b) Raoult's law and relation with molecular weight of solute.

a) కణాధార ధర్మాలు ఉదాహరణలు. b) రౌల్ట్ నియమం మరియు ద్రావిత అణుభారానికి గల

సంబంధం ను ఉత్పాదించండి.

OR

Define Osmotic pressure. Explain experimental method for determination of osmotic pressure.

ద్రవాభిసరణ పీడనం అనగా నేమి ? ద్రవాభిసరణ పీడనం ను నిర్ణయించే పద్ధతిని వివరించండి .

4. Explain Debye Huckel theory.

డిబై హుకెల్ సిద్ధాంతం ను వివరించండి

OR

Explain Conductometric titrations. వాహకత అంశ మాపనాలను వివరించండి .

5. Explain the following.

క్రింది వాటిని వివరించుము.

a) SHE హైడ్రోజన్ ఎలక్ట్రోడ్ b) Calomel electrode. కాలమ్ ఎలక్ట్రోడ్

OR

State and explain Gibbs phase rule and explain the following.
గిబ్స్ ప్రావస్థ నియమం అనగా నేమి అందలి పదాలను వివరించండి .

a) Phase b) Degrees of freedom c) Number of components.

a) ప్రావస్థ b) స్వాతంత్ర్య పరిమితులు c) ఘటకాలు సంఖ్య

SECTION – B

Answer any **FIVE** questions. Each question carries 4 marks 5 X 4 = 20 Marks
ఏవైనా ఐదు ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు నాలుగు మార్కులు

6. What are Beer-Lambert's law and its limitations?
బీర్- లాంబర్ట్ నియమం అనగానేమి ? వాటి పరిమితులు వ్రాయండి.
7. Define Chromophore and Auxochrome and give one example each?
క్రోమోఫోర్ లు మరియు ఆక్సోక్రోమోములు అనగానేమి ? ఒక్కొక్క ఉదాహరణ తో వివరించండి.
8. Write the characteristic frequencies of absorption of alcohols and aldehydes with example.
ఆల్కహాల్ మరియు ఆల్డిహైడ్ ల అభిలాక్షణిక పౌనఃపున్యం ను ఉదాహరణ తో వివరించండి.
9. Write the NMR Spectroscopy of Toluene and Ethyl acetate.
టోలెన్ మరియు ఇథైల్ ఎసిటేట్ NMR వర్ణ పటము ను వ్రాయండి.
10. Write the abnormal colligative properties with examples.
అసాధారణ కణాధార ధర్మాలను ఉదాహరణ ల తో వివరించండి.
11. Define specific conductance and Equivalent conductance.
విశిష్ట వాహకత మరియు తుల్యాంకవాహకత ను వివరించండి.
12. Write short note on potentiometric titrations.
పోటన్షియోమెట్రిక్ అంశమాపనాల పై లఘు వ్యాఖ్య వ్రాయండి.
13. Write the phase diagram of Pb-Ag system.
Pb-Ag ప్రావస్థ చిత్రం తో వివరించండి.

**Practical Paper - IV Physical Chemistry and IR Spectral Analysis
(At the end of semester IV)****Physical Chemistry****25M**

- a. Critical Solution Temperature- Phenol-Water system
- b. Effect of NaCl on critical solution temperature (Phenol-Water system)
- c. Determination of concentration of HCl conductometrically using standard NaOH solution.
- d. Determination of concentration of acetic acid conductometrically using standard NaOH Solution.

IR Spectral Analysis**25 M**

- e. IR Spectral Analysis of the following functional groups with examples
 - a) Hydroxyl groups
 - b) Carbonyl groups
 - c) Amino groups
 - d) Aromatic groups**

Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited A Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- V, Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

SEMESTER-V

Paper - V (INORGANIC, PHYSICAL & ORGANIC CHEMISTRY)

45 hrs (3 h / w)

INORGANIC CHEMISTRY

UNIT – I

Coordination Chemistry:

8h

IUPAC nomenclature-bonding theories - Review of Werner's theory and Sedgwick's concept of coordination-Valence bond theory - geometries of coordination numbers 4-tetrahedral and square planar and 6-octahedral and its limitations, crystal field theory-splitting of d-orbital's in octahedral, tetrahedral and square-planar complexes -low spin and high spin complexes - factors affecting crystal-field splitting energy, merits and demerits of crystal-field theory. Isomerism in coordination compounds - structural isomerism and stereo isomerism, stereochemistry of complexes with 4 and 6 coordination numbers.

UNIT-II

1. Spectral and magnetic properties of metal complexes:

4h

Types of magnetic behavior, spin-only formula, calculation of magnetic moments, experimental determination of magnetic susceptibility-Gouymethod.

2. Stability of metal complexes:

3h

Thermodynamic stability and kinetic stability, factors affecting the stability of metal complexes, chelate effect, determination of composition of complex by Job's method and mole ratio method.

ORGANIC CHEMISTRY

UNIT- III

Nitro hydrocarbons:

3h

Nomenclature and classification-nitro hydrocarbons, structure-Tautomerism of nitroalkanes leading to acid and keto form, Preparation of Nitroalkanes, reactivity-halogenations, reaction with HONO (Nitrous acid), Neff reaction and Mannich reaction leading to Michael addition and reduction.

UNIT- IV

Nitrogen compounds:

12h

Amines (Aliphatic and Aromatic): Nomenclature, Classification into 1°, 2°, 3° Amines and Quarternary ammonium compounds. Preparative methods -

1. Ammonolysis of alkyl halides 2. Gabriel synthesis 3. Hoffman's bromamide reaction (mechanism).

Reduction of Amides and Schmidt reaction. Physical properties and basic character - Comparative basic strength of Ammonia, methyl amine, dimethyl amine, trimethyl amine and aniline-comparative basic strength of aniline, N-methylaniline and N,N-dimethyl aniline (in aqueous and non-aqueous medium), steric effects and substituent effects. Chemical properties: a) Alkylation b) Acylation c) Carbylamine reaction d) Hinsberg separation e) Reaction with Nitrous acid of 1°, 2°, 3° (Aliphatic and aromatic amines). Electrophilic substitution of Aromatic amines -Bromination and Nitration. Oxidation of aryl and Tertiary amines, Diazotization.

PHYSICAL CHEMISTRY

UNIT- V

Thermodynamics

15h

The first law of thermodynamics-statement, definition of internal energy and enthalpy. Heat capacities and their relationship. Joule-Thomson effect- coefficient. Calculation of w , for the expansion of perfect gas under isothermal and adiabatic conditions for reversible processes. State function. Temperature dependence of enthalpy of formation-Kirchhoff's equation. Second law of thermodynamics. Different Statements of the law. Carnot cycle and its efficiency. Carnot theorem. Concept of entropy, entropy as a state function, entropy changes in reversible and irreversible processes. Entropy changes in spontaneous and equilibrium processes.

List of Reference Books

1. Concise coordination chemistry by Gopalan and Ramalingam
2. Coordination Chemistry by Basalo and Johnson
3. Organic Chemistry by G. Mare loudan, Purdue Univ
4. Advanced Physical Chemistry by
5. Text book of physical chemistry by S Glasstone
6. Concise Inorganic Chemistry by J. D. Lee
7. Advanced Inorganic Chemistry Vol-I by Satyaprakash, Tuli, Basu and Madan
8. A Text Book of Organic Chemistry by Bahl and Arun bahl
9. A Text Book of Organic chemistry by I L Finar Vol I
10. Advanced physical chemistry by Gurudeep Raj

Guidelines to the Paper Setter: The syllabus V Semester consists of Inorganic, Organic & Physical Chemistry. The V Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 08 marks from Inorganic, Organic & Physical Chemistry part.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 3 questions from Inorganic, 3 questions in Organic & 2 questions from Physical part.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT				
III B.Sc, Chemistry, V Semester				
(Inorganic, Organic & Physical chemistry)				
Sl. No	Units	Name of the chapter	8 Marks	4 Marks
INORGANIC CHEMISTRY				
1	UNIT-I	Coordination compounds	3	2
		Spectral and magnetic properties of metal complexes	1	1
ORGANIC CHEMISTRY				
2	UNIT-II	Nitro hydro carbons	3	2
		Heterocyclic compounds	1	1
PHYSICAL CHEMISTRY				
3	UNIT-III	Thermodynamics	2	2

MODEL PAPER

Dr. V.S. Krishna Govt. Degree College (A), VISAKHAPATNAM-13

III B.Sc., V SEMESTER, END EXAMINATIONS, 2019- 20

Chemistry, Paper-V

(Inorganic, Organic and Physical Chemistry-III)

(w. e. f. 2019 - 20)

Time: 3 hours

Max. Marks: 60Marks

PART- A

Answer all FIVE questions. Each question carries 8 marks. 5 x 8 = 40 Marks

అన్ని ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు ఎనిమిది మార్కులు

1. Write the important postulates of Valence Bond Theory and on the basis of VBT. How will you explain that $[\text{Ni}(\text{CO})_4]$ is Tetrahedral while $[\text{Ni}(\text{CN})_4]^{2-}$ Square planar. వేలన్ బిండ్ సిద్ధాంతం యొక్క ముఖ్య ప్రతిపాదనలు వ్రాయండి. వేలన్ బిండ్ సిద్ధాంతం అనుసరించి $[\text{Ni}(\text{CO})_4]$ టెట్రాహెడ్రల్ మరియు $[\text{Ni}(\text{CN})_4]^{2-}$ లో సమతల చదర సంశ్లేష్ట సమ్మేళనాలును వివరించండి.

(Or)

Explain crystal Field Theory. Discuss the splitting of d-orbitals in Octahedral and Tetrahedral complexes by taking any one example స్పటిక క్షేత్ర సిద్ధాంతంను వివరించండి. ఏదైనా ఒక ఉదాహరణతో అష్టముఖీయ మరియు సమతల చదర సంశ్లేష్ట సమ్మేళనాలలో-d ఆర్బిటాల్ ల విభజనను వివరించండి.?

2. How will you determine the composition of a complex by job's method and mole ration method? And write their limitations. సంశ్లేష్ట సమ్మేళనాల యొక్క సంఘటనాలను జాబ్స్ మరియు మోల్ నిష్పత్తి పద్ధతి లలో ఎట్లు నిర్ణయిస్తారు?

(Or)

Explain the absorption spectrum of $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ with suitable diagram? $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ అయాన్ శోషణ వర్ణపటమును వివరించండి?

3. (a). Explain Tautomerism in nitroalkanes with one example.

(b). Write any two methods of preparation of nitro alkanes.

©. Write any two properties of nitroalkanes.

నైట్రో అల్కేన్స్ లో టాటోమెరిజం ను ఒక ఉదాహరణ తో వివరించండి.

ఏవైనా రెండు నైట్రో అల్కేన్స్ తయారీ ధర్మాలు వ్రాయండి.

ఏవైనా రెండు నైట్రో అల్కేన్స్ రసాయన ధర్మాలు వ్రాయండి.

(Or)

(a). How do you prepare primary amines by Hoffmann Bromamide reaction method and give it's mechanism.

(b). Write the reaction of 1^0 , 2^0 , 3^0 amines with HNO_2 .

(a). ప్రైమరీ ఎమైన్లను హాఫ్ మన్ బ్రోమైడ్ చర్య పద్ధతి ద్వారా ఎట్లు తయారీ చేస్తారు. మరియు చర్య విధానమువ్రాయండి

(b). HNO_2 తో 1^0 , 2^0 , 3^0 ఎమీన్లు చర్యలను వ్రాయండి.

4. (a) How do you separate the mixture of primary, Secondary and tertiary amines by Hinsberg's method.
 (b) Write the reactions
 (i) Gabriel synthesis (ii). Carbylamines reaction
 (a). 1^0 , 2^0 , 3^0 ఎమైన్స్ ను హిన్స్ బెర్గ్ పద్ధతి ద్వారా ఏ విధముగా వేరుచేయుదురు?
 (b) క్రింది చర్యలను గురించి వివరించండి.
 (i) గాబ్రియల్ సంశ్లేషణ (ii). కార్బైల్ అమిన్ చర్య.

(Or)

- (a). Derive Kirchhoff's law?
 (b). Derive Gibbs Helmholtz equation?
 (a). కిర్కాఫ్స్ సమీకరణాన్ని ఉత్పాదించండి?
 (b) గిబ్స్ హెల్మ్ హోల్ట్ సమీకరణాన్ని ఉత్పాదించండి?
 5. Deduce the following relations
 (a). $PV^\gamma = \text{constant}$ for an adiabatic expansion of an ideal gas.
 (b). Derive $C_p - C_v = R$
 (a) ఆదర్శ వాయువులకు స్థిరోష్ణక ప్రక్రియలలో $PV^\gamma = \text{స్థిరం}$ అని నిరూపించుము?
 (b). $C_p - C_v = R$ ఉత్పాదించండి?

(Or)

- (a) Derive Carnot's cycle?
 (b). Explain the significance of entropy.
 (a) కార్నోట్ చక్రం ను ఉత్పాదించండి ?
 (b) ఎంట్రోపీ యొక్క ప్రాముఖ్యత ను వివరించండి

SECTION – B

Answer any FIVE questions. Each question carries 4 marks 5 x 4 = 20 Marks
 ఏవైనా ఐదు ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు నాలుగు మార్కులు

6. Write the postulates of Werner's theory.
 వెర్నర్ సిద్ధాంతం యొక్క ప్రతిపాదనలు వ్రాయండి
 7. What is effective atomic number and calculate the EAN of $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]$.
 ప్రాభావిక పరమాణు సంఖ్య అనగానేమి? $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]$ యొక్క ప్రాభావిక పరమాణు సంఖ్య లెక్కింపుము
 8. What is chelating effect? Give examples?
 కీలేటింగ్ ప్రభావం అనగానేమి ? ఒక ఉదాహరణ తో వివరించండి?
 9. Explain Gouy method?
 గాయి పద్ధతిని వివరించండి?
 10. Which is more basic among N, N –Dimethyl aniline and Aniline.
 N, N –డై మిథైల్ ఎనిలీన్ మరియు ఎనిలీన్ ఏది ఎక్కువ క్షార స్వభావంను ప్రదర్శించెను?
 11. Explain the basic characters of amines with the suitable examples ?
 ఎమిన్స్ యొక్క క్షార స్వభావంను సరైన ఉదాహరణ లతో వివరించండి?

12. Write any one definitions of second law of thermodynamics.
ఏదైనా ఒక ఉష్ణగతికశాస్త్ర రెండవ నియమం యొక్క నిర్వచనంను వ్రాయండి?
13. Write a short note on Joule - Thomson effect.
జౌల్-థామ్సన్ ప్రభావం ను గురించి లఘువాఖ్య వ్రాయండి?

Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited A Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- V, Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

Paper - VI (INORGANIC, ORGANIC & PHYSICAL CHEMISTRY)

45 hrs (3 h / w)

INORGANIC CHEMISTRY

UNIT-I

1. Reactivity of metal complexes:

4h

Labile and inert complexes, ligand substitution reactions - SN^1 and SN^2 , substitution reactions of square planar complexes - Trans effect and applications of Trans effect.

2. Bioinorganic chemistry:

4h

Essential elements, biological significance of Na, K, Mg, Ca, Fe, Co, Ni, Cu, Zn and Cl⁻. Metalloporphyrins – Structure and functions of hemoglobin, Myoglobin and Chlorophyll.

PHYSICAL CHEMISTRY

UNIT-II

1. Chemical kinetics

8h

Rate of reaction - Definition of order and molecularity. Derivation of rate constants for first, second, third and zero order reactions and examples. Derivation for time half change. Methods to determine the order of reactions. Effect of temperature on rate of reaction, Arrhenius equation, concept of activation energy.

2. Photochemistry

5h

Difference between thermal and photochemical processes. Laws of photochemistry- Grothus-Draper's law and Stark-Einstein's law of photochemical equivalence. Quantum yield-Photochemical reaction mechanism- hydrogen- chlorine, hydrogen- bromine reaction. Qualitative description of fluorescence, phosphorescence, and Photosensitized reactions- energy transfer processes (simple example)

ORGANIC CHEMISTRY

UNIT- III

Heterocyclic Compounds

7h

Introduction and definition: Simple five membered ring compounds with one hetero atom Ex. Furan. Thiophene and pyrrole - Aromatic character – Preparation from 1, 4, - dicarboxylic compounds, Paul-Knorr synthesis.

Properties: Acidic character of pyrrole - electrophilic substitution at 2 or 5 position, Halogenations, Nitration and Sulphonation under mild conditions - Diels Alder reaction in furan.

Pyridine- Structure - Basicity - Aromaticity - Comparison with pyrrole - one method of preparation and properties - Reactivity towards Nucleophilic substitution reaction.

UNIT-IV

Carbohydrates

8h

Monosaccharides: (+) Glucose (aldo hexose) - Evidence for cyclic structure of glucose (some negative aldehydes tests and mutarotation) - Proof for the ring size (methylation, hydrolysis and oxidation reactions) - Pyranose structure (Haworth formula and chair conformational formula).

(-) Fructose (ketohexose) - Evidence of 2 - ketohexose structure (formation of pentaacetate, formation of cyanohydrin its hydrolysis and reduction by HI). Cyclic structure for fructose (Furanose structure and Haworth formula) - osazone formation from glucose and fructose – Definition of anomers with examples.

Interconversion of Monosaccharides: Aldopentose to Aldohexose (Arabinose to D- Glucose, D-Mannose) (Kiliani - Fischer method). Epimers, Epimerisation - Lobry de bruyn van Ekenstein rearrangement. Aldohexose to Aldopentose (D-Glucose to D- Arabinose) by Ruff degradation. Aldohexose to Ketohexose [(+) Glucose to (-) Fructose] and Ketohexose to Aldohexose (Fructose to Glucose)

UNIT- V

Amino acids and proteins

7h

Introduction: Definition of Amino acids, classification of Amino acids into alpha, beta, and gamma amino acids. Natural and essential amino acids - definition and examples, classification of alpha amino acids into acidic, basic and neutral amino acids with examples. Methods of synthesis: General methods of synthesis of alpha amino acids (specific examples - Glycine, Alanine, valine and leucine) by following methods: a) from halogenated carboxylic acid b) Malonic ester synthesis c) strecker's synthesis.

Physical properties: Zwitter ion structure - salt like character - solubility, melting points, amphoteric character, definition of isoelectric point.

Chemical properties: General reactions due to amino and carboxyl groups - lactams from gamma and delta amino acids by heating peptide bond (amide linkage). Structure and nomenclature of peptides and proteins.

List of Reference Books

1. Concise coordination chemistry by Gopalan and Ramalingam
2. Coordination Chemistry by Basalo and Johnson
3. Organic Chemistry by G. Mare loudan, Purdue Univ
4. Advanced Physical Chemistry by Atkins
5. Text book of physical chemistry by S Glasstone
7. Instrumentation and Techniques by Chatwal and Anand
8. Essentials of Nano chemistry by pradeep
9. A Textbook of Physical Chemistry by Puri and Sharma
10. Advanced physical chemistry by Gurudeep Raj

LABORATORY COURSE – V
Practical Paper-V Organic Chemistry
(at the end of semester V) 30 hrs (2 h / W)

Organic Qualitative Analysis:

50M

Analysis of an organic compound through systematic qualitative procedure for functional group identification including the determination of melting point and boiling point with suitable derivatives.

Alcohols, Phenols, Aldehydes, Ketones, Carboxylic acids, Aromatic Primary Amines, Amides and Simple sugars.

LABORATORY COURSE – VI
Practical Paper – VI Physical Chemistry
(at the end of semester V) 30 hrs (2 h/W)

1. Determination of rate constant for acid catalyzed ester hydrolysis.
2. Determination of molecular status and partition coefficient of benzoic acid in Benzene and water.
3. Determination of Surface tension of liquid
4. Determination of Viscosity of liquid.
5. Adsorption of acetic acid on animal charcoal, verification of Freundlich isotherm.

Guidelines to the Paper Setter: The syllabus V Semester consists of Inorganic, Organic & Physical Chemistry. The V Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 08 marks from Inorganic, Organic & Physical Chemistry part.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 3 questions from Inorganic, 3 questions in Organic & 2 questions from Physical part.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT				
III B.Sc, Chemistry, V Semester, Paper-VI				
(Inorganic, Organic & Physical chemistry)				
Sl. No	Units	Name of the chapter	8 Marks	4 Marks
INORGANIC CHEMISTRY				
1	UNIT-I	Reactivity of Metal Complex	2	1
		Bio Inorganic Chemistry	1	0
ORGANIC CHEMISTRY				
2	UNIT-II	Heterocyclic compounds	1	2
		Carbohydrates	2	1
		Amino Acids & Proteins	1	2
PHYSICAL CHEMISTRY				
3	UNIT-III	Chemical Kinetics	2	1
		Photo Chemistry	1	1

MODEL PAPER
Dr. V.S. Krishna Govt. Degree College (A), Visakhapatnam-13
III B.Sc., V SEMESTER - END EXAMINATIONS 2019 - 20
Paper-VI, CHEMISTRY
(Inorganic, Organic and Physical Chemistry-VI)
(w. e. f. 2019 – 20)

Time: 3 hours

Max. Marks: 60Marks

PART- A

Answer all FIVE questions. Each question carries 8 marks. 5 X 8 = 40 Marks

అన్ని ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు ఎనిమిది మార్కులు

1. Explain SN1 & SN² reactions in metal complex compounds?
లోహ సంక్లిష్ట సమ్మేళనాలలో గల SN¹ & SN² చర్యలను వివరించండి?

(Or)

Discuss the Trans effect and write the theories of trans effect and its applications?
ట్రాన్స్ ప్రభావం ను వివరించి, దాని యొక్క సిద్ధాంతాలను మరియు అనువర్తనాలను వ్రాయండి?

2. Explain the structure and functions of hemoglobin, chlorophyll?

హిమోగ్లోబిన్ మరియు క్లోరోఫిల్ నిర్మాణాలను వివరించి వాటి యొక్క విధులను తెల్పండి?

(Or)

Write the classification of amino acids and write any two methods of preparations amino acids?

ఎమినో ఆమ్ల వర్గీకరణ వ్రాసి మరియు వాటి యొక్క రెండు తయారీ పద్ధతులను వ్రాయండి?

3. Define first order reaction? Derive the rate equation for first order reaction?

ప్రథమ క్రమాంక చర్య అనగానేమి? దాని యొక్క రేటు సమీకరణమను ఉత్పాదించండి?

(Or)

Write any three methods of determination of order of reaction?
చర్య క్రమాంకము నిర్ధారించు ఏవైనా మూడు పద్ధతులను గురించి వ్రాయండి?

4. What are fluorescence and phosphorescence? Explain with Joblonski diagram?

ఫ్లోరోసెన్స్ మరియు ఫాస్ఫోరిసెన్స్ అనగానేమి? జబ్లొన్స్కి పటం ద్వారా వివరించండి?

(Or)

Explain the aromatic character of furan, pyrrole, and thiophene and write the preparation methods for the above compounds?

ఫ్యూరాన్, పిర్రోల్, మరియు థయోఫిన్ ల యొక్క ఆరోమటిక్ స్వభావం గురించి వ్రాసి మరియు వాటి యొక్క తయారీ ధర్మాలను వ్రాయండి?

5. Explain the structural elucidation of glucose?

గ్లూకోజు యొక్క నిర్మాణం సోదాహరణం గా వివరించండి ?

(Or)

Explain the following క్రింది వాటిని గురించి వివరించండి?

- (i) Aldopentose to aldohexose)i) అల్టోపెంటోజ్ నుండి అల్టోహెక్సోజ్
(ii) Aldohexose to aldopentose)ii) అల్టోహెక్సోజ్ నుండి అల్టోపెంటోజ్

SECTION – B

Answer any FIVE questions. Each question carries 4 marks 5 X 4 = 20 Marks

ఏవైనా ఐదు ప్రశ్నలకు సమాధానాలు రాయండి. ప్రతి ప్రశ్నకు నాలుగు మార్కులు

6. What are labile and inert complexes? Give examples?

ఆస్థిరత్వ మరియు జడత్వ సంశ్లేష్ట సమ్మేళనాలు అనగానేమి ? ఉదాహరణ తో వివరించండి?

7. Define order and molecularity with examples?

క్రమాంకం మరియు అణుత అనగానేమి ? ఉదాహరణ తో వివరించండి?

8. State Stark-Einstein law and Grothoths draper law?

స్టార్క్ ఐనిస్టైన్ మరియు గ్రోత్స్ డ్రాపర్ నియమాలను గురించి వ్రాయండి?

9. Write chichibbin reaction with their mechanism?

చిచిబాబిన్ చర్యను చర్యా విధానం తో వివరించండి?

10. The electrophonic substitution reaction occurs in five membered hetero cyclic compounds at position-2. Explain reason?

ఎలక్ట్రోఫిలిక్ ప్రతిక్షేపన చర్యలు 5 వలయలు కలిగిన విజాతీయ వలయ సమ్మేళనాలలో 2-వ స్థానంలో మాత్రమే జరుగును. ఎందుకు ?

11. What is mutarotation? Write note in glucose mutarotation?

క్షీణభ్రమణత అనగానేమి? గ్లూకోజులో క్షీణభ్రమణతను గురించి వ్రాయండి ?

12. Define Zwitterions and Isoelectric point? Give its uses?

జ్వెట్టర్ ఆయన్ మరియు ఐసోఎలక్ట్రిక్ స్థానం లను నిర్వచించండి. వాటియొక్క ఉపయోగాలను వ్రాయండి?

13. What are peptide bonds? What are essential amino acids? Give examples?

పెప్టిడ్ బంధాలు అనగానేమి? ఆవశ్యక ఎమిన్ ఆమ్లులు అనగానేమి? వాటియొక్క ఉపయోగాలను వ్రాయండి?

DEPARTMENT OF CHEMISTRY

III B.Sc. Chemistry Cluster Electives Syllabus under CBCS
Syllabus, Blue Print & Model Question Papers

WITH EFFECT FROM 2019-20



Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited A Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Board Of Studies October- 2018

Dr. V. S. Krishna Govt. Degree & P.G. College (A), Visakhapatnam
III B.Sc. Semester-VI, Chemistry Cluster Electives Syllabus under CBCS
[w.e.f. 2019-20 (Revised in October 2018)]

Year	Semester	Paper	Title	Marks	Credits
	* Any one Paper from VII A, B and C	VII (A)*	Elective	100	03
			Practical - VII A	50	02
		VII (B)*	Elective	100	03
			Practical - VII B	50	02
		VII (C)*	Elective	100	03
			Practical - VII C	50	02
	** Any one cluster from VIII, A, B and C	VIII (A)**	Cluster Electives - I :	100	03
			VIII-A-1	100	03
			VIII-A-2	100	03
			VIII-A-3	50	02
				50	02
				50	02
		VIII (B)**	Cluster Electives - II ::	100	03
			VIII-B-1	100	03
			VIII- B-2	100	03
			VIII-B-3	50	02
				50	02
				50	02
VI	VIII (C)**	Cluster Electives - III ::	100	03	
		VIII-C-1	100	03	
		VIII-C-2	100	03	
		VIII-C-3	50	02	
			50	02	
			50	02	

Dr. V. S. Krishna Govt. Degree & PG College (A)

(NAAC Reaccredited A Grade Institution & District Identified College)

CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Semester- VI (Elective) Chemistry Syllabus

(w.e.f 2019-20 revised on October-2018)

ELECTIVE Paper- VII (A): ANALYTICAL METHODS IN CHEMISTRY

45 hrs (3h / w)

UNIT-I

Quantitative analysis:

10h

- a) Importance in various fields of science, steps involved in chemical analysis. Principles of volumetric analysis. Theories of acid-base, redox, complexometric, iodometric and precipitation titrations - choice of indicators for these titrations.
- b) Principles of gravimetric analysis: precipitation, coagulation, peptization, co-precipitation, post precipitation, digestion, filtration and washing of precipitate, drying and ignition.

UNIT-II

Treatment of analytical data:

7h

Types of errors, significant figures and its importance, accuracy - methods of expressing accuracy, error analysis and minimization of errors, precision - methods of expressing precision, standard deviation and confidence limit.

UNIT-III

SEPARATION TECHNIQUES IN CHEMICAL ANALYSIS:

8h

SOLVENT EXTRACTION: Introduction, principle, techniques, factors affecting solvent extraction, Batch extraction, continuous extraction and counter current extraction. Synergism., Application - Determination of Iron (III)

ION EXCHANGE :Introduction, action of ion exchange resins, separation of inorganic mixtures, applications, Solvent extraction: Principle and process,

UNIT – IV

10h

Chromatography: Classification of chromatography methods, principles of differential migration adsorption phenomenon, Nature of adsorbents, solvent systems, R_f values, factors effecting R_f values.

Paper Chromatography: Principles, R_f values, experimental procedures, choice of paper and solvent systems, developments of chromatogram - ascending, descending and radial. Two dimensional chromatography, applications.

UNIT -V**10h**

Thin layer Chromatography (TLC): Advantages. Principles, factors effecting R_f values. Experimental procedures. Adsorbents and solvents. Preparation of plates. Development of the chromatogram. Detection of the spots. Applications.

Column Chromatography: Principles, experimental procedures, Stationary and mobile Phases, Separation technique. Applications

HPLC : Basic principles and applications.

List of Reference Books

1. Analytical Chemistry by Skoog and Miller
2. A textbook of qualitative inorganic analysis by A.I. Vogel
3. Nanochemistry by Geoffrey Ozin and Andre Arsenault
4. Stereochemistry by D. Nasipuri
5. Organic Chemistry by Clayden

LABORATORY COURSE – VI**Practical Paper – VII-(A) (at the end of semester VI) 30hrs (2 h / W)****50M**

1. Identification of amino acids by paper chromatography.
2. Determination of Zn using EDTA
3. Determination of Mg using EDTA

Guidelines to the Paper Setter: The syllabus VI Semester consists of **Analytical methods in chemistry**. The VI Semester question paper consists of 2 sections.

In Section- A: Consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from **Analytical methods in chemistry**.

In Section- B: Consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 8 questions from **Analytical methods in chemistry**.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT

B. Sc Chemistry, Paper-VII (A), Semester- VI (ANALYTICAL METHODS IN CHEMISTRY)

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	1
3	UNIT III	1	1
4	UNIT IV	1	2
5	UNIT V	1	2

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)
B. Sc, DEGREE FINAL YEAR EXAMINATIONS
(w.e.f 2019-20 revised on October-2018)

Paper –VII (A), SEMESTER- VI
Elective: Analytical Methods in Chemistry

Time: 3 hours

Max Marks: 60M

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain about theories acid base titrations.
Or
Explain about gravimetric analysis.
2. Explain about Errors.
Or
Explain methods of expressing of precision and accuracy.
3. Explain the principle and techniques of solvent extraction.
Or
Explain the ion exchange resins.
4. Explain the principle and classification of chromatography methods.
Or
Explain about paper chromatography.
5. Explain about Thin layer chromatography.
Or
Explain about column chromatography.

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain about complexometric titrations.
7. Explain about digestion and ignition.
8. Explain about significant figures and its importance.
9. Explain determination of Fe (III) by using solvent extraction.
10. Explain the R_f value and factors effecting R_f value.
11. Explain about two dimensional chromatography and applications.
12. Explain principle and applications of HPLC.
13. Explain applications of column chromatography.

SEMESTER-VI

ELECTIVE PAPER - VII (B): ENVIRONMENTAL CHEMISTRY

45 hrs (3 h / w)

UNIT-I

Introduction

9h

Concept of Environmental chemistry-Scope and importance of environment in now a days – Nomenclature of environmental chemistry – Segments of environment - Natural resources – Renewable Resources – Solar and biomass energy and Nonrenewable resources – Thermal power and atomic energy – Reactions of atmospheric oxygen and Hydrological cycle.

UNIT-II

Air Pollution

9h

Definition-Sources of air pollution – Classification of air pollution-Acid rain-Photochemical smog-Green house effect – Formation and depletion of ozone-Bhopal gas disaster-Controlling methods of air pollution.

UNIT-III

Water pollution

9h

Unique physical and chemical properties of water – water quality and criteria for finding of water quality – Dissolved oxygen – BOD, COD, Suspended solids, total dissolved solids, alkalinity – Hardness of water – Methods to convert temporary hard water into soft water – Methods to convert permanent hard water into soft water – eutrophication and its effects – principal wastage treatment – Industrial waste water treatment.

UNIT-IV

Chemical Toxicology

9h

Toxic chemicals in the environment – effects of toxic chemicals – cyanide and its toxic effects – pesticides and its biochemical effects – toxicity of lead, mercury, arsenic and cadmium.

UNIT-V

9h

Ecosystem and biodiversity

Ecosystem

Concepts – structure – Functions and types of ecosystem – Abiotic and biotic components – Energy flow and Energy dynamics of ecosystem – Food chains – Food web – Tropic levels – Biogeochemical cycles (carbon, nitrogen and phosphorus)

Biodiversity

Definition – level and types of biodiversity – concept - significance – magnitude and distribution of biodiversity – trends - biogeographical classification of india – biodiversity at national, global and regional level.

List of Reference books

1. Fundamentals of ecology by M. C. Dash
2. A Text book of Environmental chemistry by W. Moore and F.A. Moore
3. Environmental Chemistry by Shamir K. Banerji

LABORATORY COURSE – VI

Practical Paper – Elective VII B (at the end of semester VI) 30 hrs (2 h / W)

1. Determination of carbonate and bicarbonate in water samples (acidity and alkalinity)
2. Determination of hardness of water using EDTA
 - a) Permanent hardness
 - b) Temporary hardness
3. Determination of Acidity
4. Determination of Alkalinity
5. Determination of chlorides in water samples

Guidelines to the Paper Setter: The syllabus VI Semester consists of **Environmental chemistry**. The VI Semester question paper consists of 2 sections.

In Section- A: Consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from **Environmental chemistry**.

In Section- B: Consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 8 questions from **Environmental chemistry**.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT

B. Sc Chemistry, Paper-VII (B), Semester-VI (Environmental chemistry)

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	2
3	UNIT III	1	1
4	UNIT IV	1	1
5	UNIT V	1	2

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)
B. Sc, DEGREE FINAL YEAR EXAMINATIONS
(w.e.f 2019-20 revised on October-2018)

Paper –VII (B), SEMESTER- VI
Elective: Environmental Chemistry,

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain about importance and nomenclature of environmental chemistry.
Or
Explain about atmospheric oxygen and Hydrological cycle.
2. Explain about air pollution.
Or
Explain the following
(a) Acid rain (b) Green house effect
3. Explain physical and chemical properties of water.
Or
Explain about the methods to convert permanent hard water into soft water.
4. Explain about toxicity effects of lead and mercury.
Or
Explain about toxicity effects of cyanide and pesticides.
5. Explain about bio chemical cycle of nitrogen and phosphorus.
Or
Explain about Biodiversity

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain about renewable and non renewable resources.
7. Explain about segments of environment.
8. Explain about formation and depletion of ozone.
9. Explain about controlling methods of air pollution.
10. Explain about industrial waste water treatment.
11. Explain the toxicity effect of arsenic and cadmium.
12. Explain about functions and types of Eco system.
13. Write short notes on food chain and food web.

SEMESTER-VI
ELECTIVE PAPER – VII-(C) GREEN CHEMISTRY
45 hrs (3 h / w)

UNIT-I **10h**
Green Chemistry: Introduction- Definition of green Chemistry, need of green chemistry, basic principles of green chemistry. Green synthesis- Evaluation of the type of the reaction
i) Rearrangements (100% atom economic), ii) Addition reaction (100% atom economic).
Organic reactions by Sonication method: apparatus required examples of sonochemical reactions (Heck, Hundsdiecker and Wittig reactions).

UNIT-II **10h**
Selection of solvent:i) Aqueous phase reactions ii) Reactions in ionic liquids, Heckreaction, Suzuki reactions, epoxidation. iii) Solid supported synthesis
Super critical CO₂: Preparation, properties and applications, (decaffeination, dry cleaning)

UNIT-III **10h**
Microwave and Ultrasound assisted green synthesis: Apparatus required, examples of MAOS (synthesis of fused anthro quinones, Leukart reductive amination of ketones) - Advantages and disadvantages of MAOS. Aldol condensation-Cannizzaro reaction- Diels-Alder reactions-Strecker's synthesis

UNIT-IV **5h**
Green catalysis: Heterogeneous catalysis, use of zeolites, silica, alumina, supported catalysis- biocatalysis: Enzymes, microbes Phase transfer catalysis (micellar/surfactant)

UNIT V **10h**
Examples of green synthesis / reactions and some real world cases: 1. Green synthesis of the following compounds: adipic acid, catechol, disodium imino di acetate (alternative Strecker's synthesis) 2. Microwave assisted reaction in water – Hoffmann elimination – methyl benzoate to benzoic acid–oxidation of toluene and alcohols – microwave assisted reactions in organic solvents. Diels-Alder reactions and decarboxylation reaction. 3. Ultrasound assisted reactions–sonochemical Simmons –Smith reaction(ultrasonic alternative to iodine)

Reference books:

1. Green Chemistry Theory and Practice. P.T.Anatas and J.C. Warner
2. Green Chemistry V.K. Ahluwalia Narosa, New Delhi.
3. Real world cases in Green Chemistry M.C. Cann and M.E. Connelly
4. Green Chemistry: Introductory Text M.Lancaster: Royal Society of Chemistry (London)
5. Green Chemistry: Introductory Text, M.Lancaster
6. Principles and practice of heterogeneous catalysis, Thomas J.M., Thomas M.J., John Wiley
7. Green Chemistry: Environmental friendly alternatives R S Sanghli and M.M Srivastava, Narosa Publications

LABORATORY COURSE – VII**Practical Paper – Elective VII C (at the end of semester VI) 30 hrs (2 h/W)**

1. Determination of specific reaction rate of hydrolysis for methyl acetate catalysed by hydrogen ion at room temperature.
2. Determination of molecular status and partition coefficient of benzoic acid in Benzene and water.
3. Surface tension and viscosity of liquids.
4. Adsorption of acetic acid on animal charcoal, verification of Freundlich isotherm.

Guidelines to the Paper Setter: The syllabus VI Semester consists of Green **chemistry**. The VI Semester question paper consists of 2 sections.

In Section-A: Consists of FIVE internal choice essay questions are to be set, each question carries 8 marks .The examiner has to choose 2 question from each unit from **Green chemistry**.

In Section- B: Consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered. The examiner has to choose 8 questions from **Green chemistry**.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT

B. Sc Chemistry, Paper-VII(C), Semester-VI (Green chemistry)

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	1
3	UNIT III	1	2
4	UNIT IV	1	1
5	UNIT V	1	2

MODEL PAPER

Dr. V. S Krishna Govt. Degree College, Visakhapatnam (A)

B. Sc, DEGREE FINAL YEAR EXAMINATIONS

(w.e.f 2019-20 revised on October-2018)

Paper-VII (C), SEMESTER-VI

Elective: Green Chemistry

SECTION A

Answer all questions

5 x 8 = 40M

1. Define green chemistry? Explain basic principles of green chemistry.

Or

Explain the following organic reactions by sonication method.
Heck reaction (b) Hundsdiecker reaction (c) Wittig reaction

2. Explain aqueous phase reactions and reactions in ionic solids.

Or

Explain about preparation and properties of super critical CO₂.

3. Explain about apparatus and examples of MAOS.

Or

Explain the following reactions by using MAOS.

(a) Aldol condensation (b) Diels Alder reactions (c) Strecker's synthesis

4. Explain about Heterogeneous catalysis

Or

Explain about phase transfer catalysis.

5. Explain the green synthesis of following compounds.

(a) Adipic acid (b) catechol (c) Disodium imino di acetate

Or

Explain the microwave assisted reactions in organic solvents.

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain about atom economy with examples.
7. Explain about Green synthesis.
8. Explain about solid support synthesis.
9. Explain about advantages and disadvantages of MAOS.
10. Explain about Cannizzaro's reaction by using MAOS.
11. Explain about Bio catalysis.
12. Explain the microwave assisted reaction in water.
13. Explain the Ultra sound assisted reactions.

CLUSTER ELECTIVES: Cluster Elective -I

Analytical and Physical

SEMESTER-VI

PAPER – VIII-A1: POLYMER CHEMISTRY

45 hrs (3 h / w)

UNIT-I

12h

Introduction of polymers:

Basic definitions, degree of polymerization ,classification of polymers- Natural and Synthetic polymers, Organic and Inorganic polymers, Thermoplastic and Thermosetting polymers, Plastics, Elastomers , Fibers and Resins, Linear ,Branched and Cross Linked polymers, Addition polymers and Condensation Polymers, mechanism of polymerization. Free radical, ionic and Zeigler – Natta polymerization.

UNIT-II

10h

Techniques of Polymerization: Bulk polymerization, solution polymerization, suspension and Emulsion polymerization.

Molecular weights of polymers: Number average and weight average molecular weights
Determination of molecular weight of polymers by Viscometry , Osmometry and light scattering methods.

UNIT-III

6h

Kinetics of Free radical polymerization, Glass Transition temperature(Tg) and Determination of Tg:

Free volume theory, WLF equation, factors affecting glass transition temperature (Tg).

UNIT-IV

9h

Polymer additives:

Introduction to plastic additives-fillers, Plasticizers and Softeners, Lubricants and Flow Promoters, Anti aging additives, Flame Retardants, Colorants, Blowing agents, Cross linking agents, Photo stabilizers, Nucleating agents.

UNIT-V

8h

Polymers and their applications:

Preparation and industrial applications of Polyethylene, Polyvinyl chloride, Teflon, Polyacrylonitrile, Terelene, Nylon6.6 silicones.

Reference Books:

1. Seymour, R.B. & Carraher, C.E. *Polymer Chemistry: An Introduction*, Marcel Dekker, Inc. New York, 1981.
2. Odian, G. *Principles of Polymerization*, 4th Ed. Wiley, 2004.
3. Billmeyer, F.W. *Textbook of Polymer Science*, 2nd Ed. Wiley Interscience, 1971.
4. Ghosh, P. *Polymer Science & Technology*, Tata McGraw-Hill Education, 1991.34
5. Lenz, R.W. *Organic Chemistry of Synthetic High Polymers*. Interscience Publishers, NewYork, 1967.

Guidelines to the Paper Setter: The syllabus VI Semester consists of **POLYMER CHEMISTRY**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **POLYMER CHEMISTRY**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT

**CLUSTER ELECTIVES: Cluster Elective I
Analytical and physical
SEMESTER VI
PAPER VIII A1: POLYMER CHEMISTRY**

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	2
3	UNIT III	1	1
4	UNIT IV	1	2
5	UNIT V	1	1

MODEL PAPER

Dr. V. S. Krishna Govt. Degree & PG College (A)

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Visakhapatnam- 530 013, Andhra Pradesh, INDIA

(w.e.f 2019-20 revised on October-2018)

CLUSTER ELECTIVES: Cluster Elective I, Analytical and physical

SEMESTER VI

PAPER VIII A1: POLYMER CHEMISTRY

SECTION A

Answer all questions:

5X8=40M

1. Define a polymer? Explain the classification of polymers?
Or
Explain free radical and ionic polymerization mechanism of polymers with suitable examples?
2. Explain any one method to determine the molecular weight polymers?
Or
Explain different techniques of polymerization?
3. Define Glass transition temperature? Explain the method to determine Glass transition temperature?
Or
Explain the kinetics of free radical polymerization?
4. Explain about fillers plasticizers?
Or
Explain about anti aging additives and flame retardants?
5. Explain the preparation and industrial application of poly ethylene and poly vinyl chloride?
Or
Explain the preparation and industrial application of Teflon and polyacrylonitrile?

SECTION B

Answer any five questions:

5X4=20M

6. Explain about Zeigler Natta polymerization.
7. Explain about plastics and Elastomers?
8. Explain Number average and Weight average molecular weight of polymers.
9. Explain about bulk polymerization?
10. Explain the factors affecting glass transition temperature?
11. Explain about lubricants?
12. Explain about cross linking agents?
13. Explain preparation and industrial application of Terelene?

SEMESTER-VI

PAPER – VIII-A-2: INSTRUMENTAL METHODS OF ANALYSIS

45 hrs (3 h / w)

UNIT – I

Introduction to spectroscopic methods of analysis: 4 h

Recap of the spectroscopic methods covered in detail in the core chemistry syllabus: Treatment of analytical data, including error analysis. Classification of analytical methods and the types of instrumental methods. Consideration of electromagnetic radiation.

UNIT – II

Molecular spectroscopy: 8h

Infrared spectroscopy:

Interactions with molecules: absorption and scattering. Means of excitation (light sources), separation of spectrum (wavelength dispersion, time resolution), detection of the signal (heat, differential detection), interpretation of spectrum (qualitative, mixtures, resolution), advantages of Fourier Transform (FTIR). Samples and results expected. Applications: Issues of quality assurance and quality control, Special problems for portable instrumentation and rapid detection.

UNIT – III

10h

UV-Visible/ Near IR – emission, absorption, fluorescence and photoacoustic. Excitation sources (lasers, time resolution), wavelength dispersion (gratings, prisms, interference filters, laser, placement of sample relative to dispersion, resolution), Detection of signal (photocells, photomultipliers, diode arrays, sensitivity and S/N), Single and Double Beam instruments, Interpretation (quantification, mixtures, absorption vs. fluorescence and the use of time, photoacoustic, fluorescent tags).

UNIT – IV

Separation techniques

Chromatography: Gas chromatography, liquid chromatography, supercritical fluids, Importance of column technology (packing, capillaries), Separation based on increasing number of factors (volatility, solubility, interactions with stationary phase, size, electrical field), Detection: simple vs. specific (gas and liquid), Detection as a means of further analysis (use of tags and coupling to IR and MS), Electrophoresis (plates and capillary) and use with DNA analysis. 46 *Immunoassays and DNA techniques* **8h**

Mass spectroscopy: Making the gaseous molecule into an ion (electron impact, chemical ionization), Making liquids and solids into ions (electrospray, electrical discharge, laser desorption, fast atom bombardment), Separation of ions on basis of mass to charge ratio, Magnetic, Time of flight, Electric quadrupole. Resolution, time and multiple separations, Detection and interpretation (how this is linked to excitation). **8h**

UNIT – V

Elemental analysis: **10h**

Mass spectrometry (electrical discharges).

Atomic spectroscopy: Atomic absorption, Atomic emission, and Atomic fluorescence. Excitation and getting sample into gas phase (flames, electrical discharges, plasmas), Wavelength separation and resolution (dependence on technique), Detection of radiation (simultaneous/scanning, signal noise), Interpretation (errors due to molecular and ionic species, matrix effects, other interferences).

NMR spectroscopy: Principle, Instrumentation, Factors affecting chemical shift, Spin coupling, Applications. **4h**

Electroanalytical Methods: Potentiometry & Voltammetry **4h**

Radiochemical Methods **4h**

X-ray analysis and electron spectroscopy (surface analysis)

Reference books:

1. Skoog, D.A. Holler F.J. & Nieman, T.A. *Principles of Instrumental Analysis*, Cengage Learning India Ed.
2. Willard, H.H., Merritt, L.L., Dean, J. & Settoe, F.A. *Instrumental Methods of Analysis*, 7th Ed. Wadsworth Publishing Company Ltd., Belmont, California, USA, 1988.
3. P.W. Atkins: *Physical Chemistry*.
4. G.W. Castellan: *Physical Chemistry*.
5. C.N. Banwell: *Fundamentals of Molecular Spectroscopy*.
6. Brian Smith: *Infrared Spectral Interpretations: A Systematic Approach*.
7. W.J. Moore: *Physical Chemistry*

Guidelines to the Paper Setter: The syllabus VI Semester consists of **INSTRUMENTAL METHODS OF ANALYSIS**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **INSTRUMENTAL METHODS OF ANALYSIS**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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SEMESTER VI, PAPER VIII A2

Instrumental methods of analysis

Sl. No.	UNIT	8 Marks	4 Marks
1	UNIT I	1	1
2	UNIT II	1	1
3	UNIT III	1	2
4	UNIT IV	1	2
5	UNIT V	1	2

MODEL PAPER

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Visakhapatnam- 530 013, Andhra Pradesh, INDIA

(w.e.f 2019-20 revised on October-2018)

Semester VI, Paper VIII- A2 (Instrumental methods of analysis)

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain the classification of analytical methods?
Or
Explain the classification of instrumental methods?
2. Explain the advantages of Fourier transform infrared spectroscopy?
Or
Explain the principle and different modes of vibration in infrared spectroscopy?
3. Explain about single and double beam spectrophotometers?
Or
Explain about fluorescence and photoacoustic?
4. Explain gas chromatography?
Or
Explain the principle of mass spectroscopy?
5. Explain the principle of NMR spectroscopy?
Or
Explain about atomic emission and atomic fluorescence?

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain about error analysis?
7. Explain the issues of quality assurance and quality control in FTIR?
8. Explain about photomultipliers?
9. Explain about excitation sources in UV spectroscopy?
10. Explain about Electrophoresis?
11. Write short notes on Electric quadrupole?
12. Explain about chemical shift?
13. Explain about Radio chemical methods?

SEMESTER-VI

PAPER – VIII-A-3: ANALYSIS OF DRUGS, FOODS, DAIRY PRODUCTS & BIO-CHEMICAL ANALYSIS 45 hrs (3 h / w)

UNIT- I

Analysis of the following drugs and pharmaceuticals preparations:

(Knowledge of molecular formula, structure and analysis)

Analysis of analgesics and antipyretics like aspirin

Analysis of antimalarials like chloroquine .

Analysis of drugs in the treatment of infections and infestations: penicillin.

Anti tuberculous drug- isoniazid.

UNIT - II

Analysis of the following drugs and pharmaceuticals preparations:

(Knowledge of molecular formula, structure and analysis)

Analysis of antihistamine drugs and sedatives like: allegra, zyrtec(citirizine), diazepam,

UNIT - III

Analysis of anti epileptic and anti convulsant drugs like phenobarbital.

Analysis of drugs used in case of cardiovascular drugs:atenolol.

Analysis of lipitor(atorvastatin) a drug for the prevention of production of cholesterol.

Analysis of diuretics like: furosemide (Lasix).

Analysis of prevacid(lansoprazole) a drug used for the prevention of production of acids in stomach.

UNIT - IV

Analysis of Milk and milk products: Acidity, total solids, fat, total nitrogen, proteins, lactose, phosphate activity, casein, chloride. Analysis of food materials- Preservatives: Sodium carbonate, sodium benzoate, Coloring matters,- Brilliant blue FCF. Flavoring agents - Vanilla , diacetyl, isoamyl acetate, limonene, ethylpropionate , allyl hexanoate and Adulterants in rice and wheat, wheat flour, sago, coconut oil, coffee powder, tea powder, milk..

UNIT - V

Clinical analysis of blood:Composition of blood,clinical analysis,trace elements in the body.Estimation of blood cholesterol,glucose,enzymes,RBC & WBC ,Blood gas analyser.

REFERENCE BOOKS :

- 1.F.J.Welcher-Standard methods of analysis,
- 2.A.I.Vogel-A text book of quantitative Inorganic analysis-ELBS,
- 3.F.D.Snell & F.M.Biffen-Commercial methods of analysis-D.B.Taraporavala & sons,
- 4.J.J.Elving and I.M.Kolthoff- Chemical analysis - A series of monographs on analytical chemistry and its applications -- Inter Science- Vol I to VII.,
- 5.Analytical Agricultural Chemistry by S.L.Chopra & J.S.Kanwar -- Kalyani Publishers

6. Quantitative analysis of drugs in pharmaceutical formulations by P.D.Sethi, CBS Publishers and Distributors, New Delhi
7. G.Ingram- Methods of organic elemental micro analysis- Chapman and Hall.,
8. H.Wincciam and Bobbles (Henry J)- Instrumental methods of analysis of food additives.,
9. H.Edward-The Chemical analysis of foods;practical treatise on the examination of food stuffs and the detection of adulterants,
10. The quantitative analysis of drugs- D.C.Garratt-Chapman & Hall.,
11. A text book of pharmaceutical analysis by K.A.Connors-Wiley-International.,
12. Comprehensive medicinal chemistry-Ed Corwin Hansch Vol 5,Pergamon Press.,

LABORATORY COURSE – VIII

Practical Paper – VIII-A-1: (at the end of semester VI) 30 hrs (2 h / W)

1. Preparation of Aspirin
2. Preparation of Paracetamol
3. Preparation of Acetanilide
4. Preparation of Barbutiric Acid
5. Preparation of Phenyl Azo β -naphthol

LABORATORY COURSE – VIII

Practical Paper – VIII-A-2 (at the end of semester VI) 30 hrs (2 h/ W)

1. Green procedure for organic qualitative analysis: Detection of N, S and halogens
2. Acetylation of 1^o amine by green method: Preparation of acetanilide
3. Rearrangement reaction in green conditions: Benzil-Benzilic acid rearrangement
4. Electrophilic aromatic substitution reaction: Nitration of phenol
5. Radical coupling reaction: Preparation of 1, 1-bis -2-naphthol
6. Green oxidation reaction: Synthesis of adipic acid
7. Green procedure for Diels Alder reaction between furan and maleic anhydride

List of Reference Books

1. Green Chemistry Theory and Practice. P.T.Anatas and J.C. Warner
2. Green Chemistry V.K. Ahluwalia Narosa, New Delhi.
3. Real world cases in Green Chemistry M.C. Cann and M.E. Connelly
4. Green Chemistry: Introductory Text M.Lancaster: Royal Society of Chemistry (London)
5. Green Chemistry: Introductory Text, M.Lancaster
6. Principles and practice of heterogeneous catalysis, Thomas J.M., Thomas M.J., John Wiley
7. Green Chemistry: Environmental friendly alternatives R S Sanghli and M.M Srivastava, Narosa Publications

Guidelines to the Paper Setter: The syllabus VI Semester consists of **ANALYSIS OF DRUGS, FOODS, DAIRY PRODUCTS & BIO CHEMICAL ANALYSIS**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **ANALYSIS OF DRUGS, FOODS, DAIRY PRODUCTS & BIO CHEMICAL ANALYSIS**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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**SEMESTER-VI, PAPER VIII- A3
ANALYSIS OF DRUGS, FOODS, DAIRY PRODUCTS
&
BIO CHEMICAL ANALYSIS**

Sl. No.	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	1
3	UNIT III	1	2
4	UNIT IV	1	2
5	UNIT V	1	1

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SEMESTER-VI, PAPER VIII- A3

Analysis of drugs, foods, dairy products & bio chemical analysis

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain the analysis of analgesics and antipyretics of Aspirin?
Or
Explain the analysis of chloroquine?
2. Explain the analysis of Allegra?
Or
Explain the analysis of Diazepam?
3. Explain the analysis of Phenobarbital?
Or
Explain the analysis of Atenolol?
4. Explain analysis the milk (Acidity, casein, fat, proteins)
Or
Explain the preservatives of sodium carbonate and sodium benzoate?
5. Explain clinical analysis of Blood?
Or
Explain the trace elements in the body

SECTION B

Answer any five questions:

5 x 4 = 20M

6. Explain the analysis of Penicillin?
7. Explain the analysis of Ionized?
8. Explain about analysis of Citirizine?
9. Explain about the analysis of Lipitor?
10. Explain the analysis of prevacid?
11. Write about flooring agents?
12. Write about food adulterants?
13. Explain about RBC and WBC?

VII-A-3 Practical:- Project Work

Cluster Elective –II

Fuels and Industrial Inorganic Materials

PAPER – VIII-B-1: FUEL CHEMISTRY AND BATTERIES

45 hrs (3 h / w)

UNIT –I

12h

Review of energy sources (renewable and non-renewable)-classification of fuels and their calorific value. Coal: Uses of Coal (fuel and non fuel) in various industries , its composition , carbonization of coal-coal gas , producer gas and water gas- composition and uses-fractionation of coal tar-uses of coal tar based chemicals , requisites of a good metallurgical coke, coal gasification (Hydro gasification and catalytic gasification) coal liquefaction and solvent refining.

UNIT-II

6h

Petroleum and petrol chemical industry:

Composition of crude petroleum, refining and different types of petroleum products and their applications.

UNIT-III

10h

Fractional distillation (principle and process) , cracking (Thermal and catalytic cracking). Reforming petroleum and non petroleum fuels (LPG , CNG , LNG , biogas) ,fuels derived from biomass , fuel from waste , synthetic fuels (gaseous and liquids) , clear fuels , petro chemicals : vinyl acetate , propylene oxide , isoprene , butadiene , toluene and its derivative xylene.

UNIT-IV

10h

Lubricants:

Classification of lubricants, lubricating oils(conducting and non conducting) , solid and semi solid lubricants , synthetic lubricants. Properties of lubricants (viscosity index , cloud point , pore point) and their determination.

UNIT-V

7h

Batteries:

Primary and secondary batteries, battery components and their role, Characteristics of Battery. Working of following batteries: Pb acid, Li-Battery, Solid state electrolyte battery. Fuel cells, Solar cell and polymer cell.

Reference books:

1. E.Stochi : Industrial chemistry , Vol-1, Ellis Horwood Ltd.UK
2. P.C.Jain , M.Jain: Engineering chemistry, Dhanpat Rai &sons , Delhi.
3. B.K.Sharma: Industrial Chemistry, Goel Publishing house, Meerut.

Guidelines to the Paper Setter: The syllabus VI Semester consists of **Fuels and industrial inorganic materials**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **Fuels and industrial inorganic materials**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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CLUSTER ELECTIVE II

Fuels and industrial inorganic materials

PAPER-VIII-B1: Fuel chemistry and batteries

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	2
3	UNIT III	1	1
4	UNIT IV	1	2
5	UNIT V	1	1

MODEL PAPER
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CLUSTER ELECTIVE-II, Fuels and industrial inorganic materials
PAPER-VIII-B1: Fuel chemistry and Batteries

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain about classification of fuels and their calorific value.

Or

Write brief note on coal tar.

2. Explain about composition of crude petroleum.

Or

Explain different types of petroleum products and their applications.

3. Explain about reforming petroleum and non petroleum fuels.

Or

Explain the fractional distillation and cracking.

4. Explain classification of lubricants.

Or

Explain about properties of lubricants.

5. Explain the primary and secondary batteries.

Or

Explain battery components and their role and characteristics of battery.

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain the uses of coal in various industries.
7. Explain about renewable and non renewable energy sources.
8. Explain refining of crude petroleum.
9. Write about synthetic fuels.
10. Write about petro chemicals.
11. Write about synthetic lubricants.
12. Explain the working of Li battery.
13. Explain the working of fuel cells.

SEMESTER-VI

PAPER-VIII-B-2: INORGANIC MATERIALS OF INDUSTRIAL IMPORTANCE

45 hrs (3 h / w)

UNIT - I

Recapitulation of *s*- and *p*-Block Elements

8h

Periodicity in *s*- and *p*-block elements with respect to electronic configuration, atomic and ionic size, ionization enthalpy, electronegativity (Pauling, Mulliken, and Alfred - Rochow scales). Allotropy in C, S, and P. Oxidation states with reference to elements in unusual and rare oxidation states like carbides and nitrides), inert pair effect, diagonal relationship and anomalous behaviour of first member of each group.

UNIT – II

15h

Silicate Industries

Glass: Glassy state and its properties, classification (silicate and non-silicate glasses).

Manufacture and processing of glass. Composition and properties of the following types of glasses: Soda lime glass, lead glass, armoured glass, safety glass, borosilicate glass, fluorosilicate, coloured glass, photosensitive glass.

Ceramics: Important clays and feldspar, ceramic, their types and manufacture. High technology ceramics and their applications, superconducting and semiconducting oxides, fullerenes carbon nanotubes and carbon fibre.

Cements: Classification of cement, ingredients and their role, Manufacture of cement and the setting process, quick setting cements.

UNIT – III

8h

Fertilizers:

Different types of fertilizers. Manufacture of the following fertilizers: Urea, ammonium nitrate, calcium ammonium nitrate, ammonium phosphates; polyphosphate, superphosphate, compound and mixed fertilizers, potassium chloride, potassium sulphate.

UNIT – IV

8h

Surface Coatings:

Objectives of coatings surfaces, preliminary treatment of surface, classification of surface coatings. Paints and pigments-formulation, composition and related properties. Oil paint, Vehicle, modified oils, Pigments, toners and lakes pigments, Fillers, Thinners, Enamels, emulsifying agents. Special paints (Heat retardant, Fire retardant, Eco-friendly paint, Plastic paint), Dyes, Wax polishing, Water and Oil paints, additives, Metallic coatings (electrolytic and electroless), metal spraying and anodizing.

UNIT – V

6h

Alloys:

Classification of alloys, ferrous and non-ferrous alloys, Specific properties of elements in alloys. Manufacture of Steel (removal of silicon decarbonization, demanganization, desulphurization dephosphorisation) and surface treatment (argon treatment, heat treatment, nitriding, carburizing). Composition and properties of different types of steels.

Chemical explosives:

Origin of explosive properties in organic compounds, preparation and explosive properties of lead azide, PETN, cyclonite (RDX). Introduction to rocket propellants.

Reference Books:

E. Stocchi: *Industrial Chemistry*, Vol-I, Ellis Horwood Ltd. UK.

R. M. Felder, R. W. Rousseau: *Elementary Principles of Chemical Processes*, Wiley Publishers, New Delhi.

W. D. Kingery, H. K. Bowen, D. R. Uhlmann: *Introduction to Ceramics*, Wiley Publishers, New Delhi.

J. A. Kent: *Riegel's Handbook of Industrial Chemistry*, CBS Publishers, New Delhi.

P. C. Jain & M. Jain: *Engineering Chemistry*, Dhanpat Rai & Sons, Delhi.

R. Gopalan, D. Venkappayya, S. Nagarajan: *Engineering Chemistry*, Vikas Publications, New Delhi.

B. K. Sharma: *Engineering Chemistry*, Goel Publishing House, Meerut

Guidelines to the Paper Setter: The syllabus VI Semester consists of **Inorganic materials of industrial importance**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **Inorganic materials of industrial importance**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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PAPER- VIII B-2:

Inorganic materials of industrial importance

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	2
3	UNIT III	1	1
4	UNIT IV	1	1
5	UNIT V	1	2

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PAPER- VIII B2, Inorganic materials of industrial importance

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain electronegativity of s and p block elements by using Pauling and Mulliken scales.

Or

Explain the allotropy of C, S and P.

2. Explain the composition and properties of the following

- (a) Soda lime glass (b) borosilicate glass
(b) Photo sensitive glass (d) colored glass

Or

Explain the classification, manufacture and setting process of cement.

3. Explain the manufacture of the following fertilizers.

- (a) Urea (b) ammonium nitrate (c) calcium ammonium nitrate

Or

Explain the manufacture of the following fertilizers

- (a) Ammonium phosphate (b) poly phosphate (c) super phosphate

4. Explain about special paints.

Or

Explain the paints and pigments.

5. Explain the classification and properties of alloys.

Or

Explain the composition and properties of different types of steels.

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain atomic size and ionization enthalpy of s and p block elements.
7. Explain about diagonal relationship of elements.
8. Write short notes on Ceramics.
9. Explain about classification of Glass.
10. Write about mixed fertilizers.
11. Explain classification of surface coatings.
12. Explain the surface treatment of steels.
13. Explain preparation and explosive Properties of RDX.

SEMESTER-VI

PAPER – VIII-B-3: ANALYSIS OF APPLIED INDUSTRIAL PRODUCTS

45 hrs (3 h / w)

UNIT-I

Analysis of soaps: moisture and volatile matter, combined alkali, total fatty matter, free alkali, total fatty acid, sodium silicate and chlorides.

Analysis of paints : Vehicle and pigments , Barium Sulphate , total lead, lead chromate, iron pigments, zinc chromate

UNIT- II

Analysis of oils: saponification value, iodine value, acid value, ester value, bromine value, acetyl value.

Analysis of industrial solvents like benzene, acetone, methanol and acetic acid., Determination of methoxyl and N-methyl groups.,

UNIT-III

Analysis of fertilizers: urea, NPK fertilizer, super phosphate,

Analysis of DDT, BHC, endrin, endosulfone, malathion, parathion.,

Analysis of starch, sugars, cellulose and paper,

UNIT -IV

Gas analysis: carbon dioxide, carbon monoxide, oxygen, hydrogen, saturated hydrocarbon, unsaturated hydrocarbons, nitrogen, octane number, cetane number

Analysis of Fuel gases like: water gas, producer gas, kerosene (oil) gas.

Ultimate analysis : carbon, hydrogen, nitrogen, oxygen, phosphorus and sulfur.,

UNIT - V

Analysis of Complex materials:

Analysis of cement- loss on ignition, insoluble residue, total silica, sesquioxides, lime, magnesia, ferric oxide, sulphuric anhydride.

Analysis of glasses - Determination of silica, sulphur, barium, arsenic, antimony, total R_2O_3 , calcium, magnesium, total alkalies, aluminium, chloride, fluoride

SUGGESTED BOOKS:

- 1.F.J.Welcher-Standard methods of analysis,
- 2.A.I.Vogel-A text book of quantitative Inorganic analysis-ELBS,
- 3.H.H.Willard and H.Deal- Advanced quantitative analysis- Van Nostrand Co,
- 4.F.D.Snell & F.M.Biffen-Commercial methods of analysis-D.B.Taraporavala & sons,
- 5.J.J.Elving and I.M.Kolthoff- Chemical analysis - A series of monographs on analytical chemistry and its applications -- Inter Science- Vol I to VII.,
- 6.G.Z.Weig - Analytical methods for pesticides, plant growth regulators and food additives - Vols I to VII,
- 7.Analytical Agricultural Chemistry by S.L.Chopra & J.S.Kanwar -- Kalyani Publishers
- 8.Manual of soil, plant, water and fertilizer analysis, R.M.Upadhyay and N.L Sharma, Kalyani Publishers

LABORATORY COURSE – VIII

Practical Paper – VIII-B-1: (at the end of semester VI)

30 hrs (2 h / W)

1. Preparation of Aspirin
2. Preparation of Paracetamol
4. Preparation of Acetanilide
5. Preparation of Barbutiric Acid
6. Preparation of Phenyl Azo β -naphthol

I. LABORATORY COURSE – VIII
Practical Paper – VIII-B-2: (at the end of semester VI)

30 hrs (2 h / W)

7. 1.Green procedure for organic qualitative analysis: Detection of N, S and halogens
8. 2.Acetylation of 1^o amine by green method: Preparation of acetanilide
3. Rearrangement reaction in green conditions: Benzil-Benzilic acid rearrangement
5. Electrophilic aromatic substitution reaction: Nitration of phenol
6. Radical coupling reaction: Preparation of 1,1-bis -2-naphthol
9. Green oxidation reaction: Synthesis of adipic acid
10. Green procedure for Diels Alder reaction between furan and maleic anhydride

List of Reference Books

1. Green Chemistry Theory and Practice. P.T.Anatas and J.C. Warner
2. Green Chemistry V.K. Ahluwalia Narosa, New Delhi.
4. Real world cases in Green Chemistry M.C. Cann and M.E. Connelly
3. Green Chemistry: Introductory Text M.Lancaster: Royal Society of Chemistry (London)
6. Green Chemistry: Introductory Text, M.Lancaster
7. Principles and practice of heterogeneous catalysis, Thomas J.M., Thomas M.J., John Wiley
4. Green Chemistry: Environmental friendly alternatives R S Sanghli and M.M Srivastava, Narosa Publications

Guidelines to the Paper Setter: The syllabus VI Semester consists of **Analysis of applied industrial products**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **Analysis of applied industrial products**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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PAPER-VIII B-3

Analysis of applied industrial products

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	1
2	UNIT II	1	1
3	UNIT III	1	2
4	UNIT IV	1	2
5	UNIT V	1	2

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PAPER-VIII B-3: Analysis of applied industrial products

SECTION A

Answer all questions:

5 x 8 =

40M

1. Explain the analysis of soap.

Or

Explain the analysis paints.

2. Explain the analysis of oils.

Or

Explain the analysis of following industrial solvents.

(a) Benzene (b) Acetone (c) methanol

3. Explain the analysis of following fertilizers.

(a) Urea (b) NPK fertilizer

Or

Explain the analysis of following fertilizers (a) DDT (b) Malathion

4. Explain the following gas analysis.

(a) Octane number (b) saturated and unsaturated hydrocarbons

Or

Explain the analysis of fuel gases.

5. Explain the following factors in analysis of cement.

(a) Lime (b) loss on ignition (c) ferric oxide (d) magnesia

Or

Explain the following factors in analysis of glasses.

(a) Sulphur (b) Barium (c) Arsenic (d) Antimony

SECTION B

Answer any five questions:

5 x 4 =

20M

6. Explain about vehicle and pigments.
7. Explain about determination of N –methyl and methoxyl groups.
8. Explain about the analysis of sugars.
9. Explain about the analysis of papers.
10. Write short note on Cetane number?
11. Explain the analysis nitrogen and oxygen.
12. Explain about determination of Silica.
13. Explain about sesqui oxides.

VII-A-3 Practical: - Project Work / Intern Ship

Cluster Elective - III

ORGANIC

PAPER – **VIII-C-1: ORGANIC SPECTROSCOPIC TECHNIQUES 45 hrs (3 h / w)**

UNIT-I

10h

NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Nuclear spin, Principles of NMR-Classical and Quantum Mechanical methods, Magnetic moment and Spin angular momentum. Larmor Frequency. Instrumentation. Relaxation-spin-spin & spin lattice relaxation. Shielding constants, Chemical shifts, Shielding and Deshielding mechanism-Factors influencing Chemical shift. Spin-Spin interactions-AX, AX₂ and AB types. Vicinal, Geminal and Long range coupling- Factors influencing coupling constants.

UNIT – II

5h

Spin decoupling, Spin tickling, Deuterium exchange, Chemical shift reagents and nuclear overhauser effect. Applications in Medical diagnostics, Reaction kinetics and mechanically induced dynamic nuclear polarization. FT NMR and its Advantages.

UNIT-III

10h

UV & VISIBLE SPECTROSCOPY

Electronic spectra of diatomic molecules. The Born-oppenheimer approximation. Vibrational coarse structure: Bond association and Bond sequence. Intensity of Vibrational-electronic spectra: The Franck-Condon principle. Rotational fine structure of electronic vibration transitions. Electronic structure of diatomic molecules.

Types of transitions, Chromophores, Conjugated dienes, trienes and polyenes, unsaturated carbonyl compounds-Woodward – Fieser rules.

UNIT-IV

5h

Electronic spectra of polyatomic molecules. Chemical analysis by Electronic Spectroscopy – Beer-Lambert's Law. Deviation from Beer's law. Quantitative determination of metal ions (Mn⁺², Fe⁺², NO₂⁻, Pb⁺²). Simultaneous determination of Chromium and Manganese in a mixture.

Electron Spin Resonance Spectroscopy

Basic Principles, Theory of ESR, Comparison of NMR & ESR. Instrumentation, Factors affecting the 'g' value, determination of 'g' value. Isotropic and Anisotropic constants. Splitting hyper fine splitting coupling constants. Line width, Zero field splitting and Kramer degeneracy. Crystal field splitting, Crystal field effects.

Applications:- Detection of free radicals; ESR spectra of (a) Methyl radical (CH_3), (b) Benzene anion (C_6H_6^-) (c) Isoquinine (d) $[\text{Cu}(\text{H}_2\text{O})_6]^{+2}$ (e) $[\text{Fe}(\text{CN})_5\text{NO}]^{-3}$ (f)

REFERENCE BOOKS:

1. Electron Spin Resonance Elementary Theory and Practical Applications- John E. Wertz and James R. Bolton, Chapman and Hall, 1986.
2. Spectroscopic Identification of organic compounds – Silverstein, Basseler and Morrill.
3. Organic Spectroscopy- William Kemp.
4. Fundamentals of Molecular Spectroscopy- C.N.Banwell and E.A. Mc cash 4th Edition, Tata Mc Graw Hill Publishing Co., Ltd. 1994.
5. Physical Methods in Inorganic Chemistry – R.S.Drago, Saunders Publications.
6. Application of Mössbauer Spectroscopy – Green Mood.
7. NMR, NQR, EPR and Mössbauer Spectroscopy in inorganic chemistry – R.V Parish, Ellis, Harwood.
8. Instrumental Methods of Chemical Analysis- H.Kaur, Pragathi Prakashan, 2003.
9. Instrumental Methods of Analysis, 7th Edition – Willard, Merrit, Dean, Settle, CBS Publications, 1986.
10. Molecular Structure and Spectroscopy – G. Aruldhas, Prentice Hall of India Pvt.Ltd, New Delhi, 2001.
11. Mössbauer Spectroscopy – N.N. Green Wood and T.C. Gibb, Chapman, and Hall, Landon 1971.
12. Coordination Chemistry: Experimental Methods- K. Burger, London Butter Worths, 1973.
13. Analytical spectroscopy – Kamlesh Bansal, Campus books, 2008.
14. Structural Inorganic Chemistry Mössbauer Spectroscopy – Bhide.
15. Principle of Mössbauer Spectroscopy – T.C. Gibb, Chapman, and Hall, Landon 1976.

Guidelines to the Paper Setter: The syllabus VI Semester consists of **Organic spectroscopic techniques**. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from **Organic spectroscopic techniques**.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

BLUE PRINT
CLUSTER ELECTIVE-III, ORGANIC
PAPER -VIII C 1: Organic spectroscopic techniques

Sl. No.	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	2
3	UNIT III	1	1
4	UNIT IV	1	1
5	UNIT V	1	2

MODEL PAPER

Dr. V. S. Krishna Govt. Degree & PG College (A)

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Visakhapatnam- 530 013, Andhra Pradesh, INDIA

(w.e.f 2019-20 revised on October-2018)

Cluster Elective-III, ORGANIC

Paper -VIII C 1: Organic spectroscopic techniques

SECTION A

Answer all questions:

5 x 8 = 40M

1. Explain the principle of NMR spectroscopy.
Or
Explain the spin-spin interactions of AX, AX₂ and AB types.
2. Explain about FT-NMR and its advantages.
Or
Explain about applications in medical diagnostics, reaction kinetics and mechanically induced dynamic nuclear polarization.
3. Explain about different types of electronic transitions.
Or
Explain about Woodward-Fisher rules with suitable examples.
4. Explain Beer- Lambert's law and its deviations.
Explain determination of Mn⁺² by using Beer's law.
Or
Explain about electronic spectra of poly atomic molecules.
5. Explain the basic principle and theory of Electronic spin resonance spectroscopy.
Or
Explain the following compounds using ESR spectroscopy.
(a) Isoquinine (b)[Cu(H₂O)₆]⁺²

SECTION B

Answer any five questions:

5 x 4 = 20M

6. Explain about chemical shift.
7. Explain factors influencing coupling constants.
8. Explain about nuclear over Hauser effect.
9. Explain about spin decoupling and spin tickling.
10. Explain about Frank Condon principle.
11. Explain simultaneous determination of chromium and manganese in a mixture.
12. Explain about Kramer degeneracy.
13. Explain hyperfine splitting and coupling constants in ESR spectroscopy.

Cluster Elective –III

ORGANIC

PAPER - VIII-C-2: ADVANCED ORGANIC REACTIONS 45 hrs (3 h / w)

UNIT – I

ORGANIC PHOTOCHEMISTRY

Organic photochemistry : Molecular orbitals, carbonyl chromophore–triplet states, Jablonski diagram, inter–system crossing. Energy transfer. Energies properties and reaction of singlet and triplet states of and transitions.

Photochemical reactions : (a) Photoreduction, mechanism, influence of temperature, solvent, nature of hydrogen donors, structure of substrates on the course of photo reduction,.

UNIT – II

ORGANIC PHOTOCHEMISTRY

Norrish cleavages, type I: Mechanism, acyclic cyclicdiones, influence of sensitizer, photo Fries rearrangement. Norrish type II cleavage: Mechanism and stereochemistry, type II reactions of esters: 1: 2 diketones, photo decarboxylation., Di - π methane rearrangement, Photochemistry – of conjugated dienes, Decomposition of nitrites – Barton reaction.

UNIT – III

PROTECTING GROUPS AND ORGANIC REACTIONS

Principles of (1) Protection of alcohols – ether formation including silyl ethers-ester formation, (2) Protection of carboxylic acids – ester formation, benzyl and t–butyl esters, (3) Protection of carbonyl groups-acetal, ketal, 1,2-glycols and 1,2-dithioglycols formation.

UNIT-IV

Synthetic reactions: Mannich reaction – Mannich bases – Robinson annulations. Stork-enamine reaction. Phase transfer catalysis-mechanisms and use of benzyl trialkyl ammonium halides. Wittig reaction.

UNIT –V: NEW SYNTHETIC REACTIONS

Baylis–Hillman reaction, RCM olefin metathesis, Grubb catalyst, Heck reaction, Suzuki coupling, Stille coupling and Sonogishira coupling, Click reaction.

Recommended Books

1. Molecular reactions and Photochemistry by Charles Dupey and O.L. Chapman.
2. Molecular Photochemistry by Turru.
3. Importance of antibonding orbitals by Jaffe and Orchin.
4. Text Book of Organic Chemistry by Cram,. Hammand and Henrickson.
5. Some modern methods of organic synthesis by W. Carruthers.
6. Guide Book to Organic Synthesis by R.K. Meckie, D.M. Smith and R.A. Atken.
7. Organic Synthesis by O.House.
8. Organic synthesis by Michael B. Smith.
9. Organic Chemistry Claydon and others 2005.
10. Name Reactions by Jie Jack Li
11. Reagents in Organic synthesis by B.P. Mundy and others.
12. Tandem Organic Reactions by Tse–Lok Ho.

Guidelines to the Paper Setter: The syllabus VI Semester consists of ADVANCED ORGANIC REACTIONS. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from ADVANCED ORGANIC REACTIONS.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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PAPER-VIII C2: ADVANCED ORGANIC REACTIONS

Sl. No	UNIT	8 Marks	4 Marks
1	UNIT I	1	2
2	UNIT II	1	1
3	UNIT III	1	1
4	UNIT IV	1	2
5	UNIT V	1	2

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PAPER-VIII C2: ADVANCED ORGANIC REACTIONS

SECTION A

Answer all questions:

5 x 8 = 40M

1. Explain about Joblonski diagram.
Or
Explain about photo reduction.
2. Explain about Norrish type I reactions.
Or
Explain about Norrish type II reactions
3. Explain about protection of alcohols with suitable examples.
Or
Explain about protection of carbonyl compounds with suitable examples.
4. Explain the phase transfer catalysis.
Or
Explain Mannich reaction with mechanism.
5. Explain about Bayils-Hillman reaction.
Or
Explain the following reactions.
(a) Suzuki coupling (b) Stille coupling

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain about reactions of singlet.
7. Explain structure of substrates on the course of photo reduction.
8. Explain the Barton reaction.
9. Explain about protection of carboxylic acids.
10. Explain about Witting reaction.
11. Explain Stork examine reaction.
12. Explain about click reaction.
13. Explain about RCM olefin metathesis.

Cluster Elective-III, ORGANIC

PAPER -VIII-C-3: PHARMACEUTICAL AND MEDICINAL CHEMISTRY45hrs (3h / w)

UNIT-I **8h**

Pharmaceutical chemistry Terminology: Pharmacy, Pharmacology, Pharmacophore, Pharmacodynamics, Pharmacokinetics (ADME, Receptors - brief treatment) Metabolites and Anti metabolites.

UNIT-II

Drugs: **8h**

Nomenclature: Chemical name, Generic name and trade names with examples
Classification: Classification based on structures and therapeutic activity with one example each, Administration of drugs

UNIT-III

Synthesis and therapeutic activity of the compounds: **12h**

a. Chemotherapeutic Drugs

1.Sulphadugs(Sulphamethoxazole) 2.Antibiotics - β -Lactam Antibiotics, Macrolide Antibiotics, 3. Anti malarial Drugs (chloroquine)

b. Psycho therapeutic Drugs:

1. Anti pyretics(Paracetamol) 2.Hypnotics, 3.Tranquilizers(Diazepam) 4.Levodopa

UNIT-IV

Pharmacodynamic Drugs: **8h**

1. Antiasthma Drugs (Solbutamol) 3. Antianginals (Glycerol Trinitrate)
4. Diuretics (Frusemide)

UNIT-V

HIV-AIDS: **9h**

Immunity - CD-4cells, CD-8cells, Retro virus, Replication in human body, Investigation available, prevention of AIDS, Drugs available - examples with structures: PIS: Indivanir (rixivan), Nelfinavir(Viracept).

List of Reference Books:

- 1.Medicinal Chemistry by Dr. B.V.Ramana
2. Synthetic Drugs by O.D.Tyagi & M.Yadav
3. Medicinal Chemistry by Ashutoshkar
4. Medicinal Chemistry by P.Parimoo
5. Pharmacology& Pharmacotherapeutics R.S Satoshkar & S.D.Bhandenkar
6. Medicinal Chemistry by Kadametal P-I & P.II
7. European Pharmacopoeia

Guidelines to the Paper Setter: The syllabus VI Semester consists of PHARMACEUTICAL AND MEDICINAL CHEMISTRY. The VI Semester question paper consists of 2 sections.

In Section- A: consists of FIVE internal choice essay questions are to be set, each question carries 8 marks. The examiner has to choose 2 questions from each unit from PHARMACEUTICAL AND MEDICINAL CHEMISTRY.

In Section- B: consists of EIGHT short answer questions carries 4 marks out of which 5 are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

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PAPER- VIII C 3: PHARMACEUTICAL AND MEDICINAL CHEMISTRY

Sl. No.	UNIT	8 Marks	4 Marks
1	UNIT I	1	1
2	UNIT II	1	1
3	UNIT III	1	3
4	UNIT IV	1	1
5	UNIT V	1	2

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PAPER- VIII C 3: Pharmaceutical and medicinal chemistry

SECTION A

Answer all questions

5 x 8 = 40M

1. Explain the following
(a) Pharmacy (b) pharmacology (c) pharmacophore
Or
Explain the following
(a) Pharmacodynamics (b) pharmacokinetics
2. Explain the classification of drugs.
Or
Explain administration of drugs.
3. Explain the synthesis and therapeutic activity of chloroquine.
Or
Explain the synthesis and therapeutic activity of antibiotics.
4. Explain the synthesis and therapeutic activity of salbutamol.
Or
Explain the synthesis and therapeutic activity of Glycerol trinitrate.
5. Explain the various drugs available in the treatment of HIV- AIDS.
Or
Define HIV- AIDS? Explain the factors that cause HIV- AIDS.

SECTION B

Answer any five questions

5 x 4 = 20M

6. Explain about metabolites and anti metabolites.
7. Explain the nomenclature of drugs.
8. Explain the synthesis and therapeutic activity of Levodopa.
9. Explain the synthesis and therapeutic activity of paracetamol.
10. Explain the synthesis and therapeutic activity of Diuretics.
11. Write the structure and therapeutic activity of sulphamethoxazole
12. Write short notes on CD₄ and CD₈ cells.
13. Explain the Retro virus.

DEPARTMENT OF CHEMISTRY

Inter Disciplinary Add On Course CLINICAL LAB TECHNOLOGY

Model Curriculum- Inter Disciplinary Add on Course
Syllabus & Model Question Papers

WITH EFFECT FROM 2018-19



Dr. V. S. Krishna Govt. Degree & PG College (A)

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CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Board Of Studies October- 2018

Dr. V. S. Krishna Govt. Degree & PG College (A)

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CENTRE FOR RESEARCH STUDIES

Visakhapatnam- 530 013, Andhra Pradesh, INDIA

Dept of Chemistry Inter Disciplinary Add On Course

Clinical Lab Technology-Syllabus

(w.e.f 2019-20 revised on October-2018)

MODULE- I (Clinical Lab Technology)	18 hrs
1. Solutions:	2 hrs
Definition, different concentration methods	
2. Carbohydrates and Lipids	4 hrs
(a) Carbohydrates: Definition, biological importance, classification, qualitative tests	
(b) Lipids: Definition, biological importance, classification and clinical importance.	
3. Amino acids Proteins	4 hrs
Definition, biological importance, classification, qualitative tests.	
4. Vitamins and Minerals	4 hrs
(a) Vitamins: Classification, source, deficiency diseases	
(b) Minerals: source, importance, deficiency diseases and clinical tests	
5. Blood composition	4 hrs
Blood composition functions tests	
6. Practical	12hrs
(a) Blood tests, group identification, hemoglobin, other tests	
(b) Blood sugar test	
(c) Lipid profile test	
(d) Liver function test	

(e) Kidney function test

(f) Serum calcium, Sodium, Potassium test

Guidelines to the Paper Setter:

The syllabus Inter Disciplinary Add on Course consists of Clinical Lab Technology. The question paper consists of 3 sections.

In Section-A: consists of 6 very short answers questions in which 5 are to be answered each question carries 2 marks.

In Section-B: consists of 8 short answer questions carries 5 marks out of which 5 are to be answered.

In section-C: 6 essay questions are to be set each question carrying 10 marks out of which 4 questions are to be answered.

The examiner is requested to set question in such a way that the entire syllabus is reflected in the question paper set by him.

Inter Disciplinary Add on Course Clinical Lab Technology BLUE PRINT

Sl. No	Name of the chapter	2 marks	5 marks	10 marks
Module-I (Clinical Lab Technology)				
1.	Solutions	1	1	--
2.	Carbohydrates and Lipids	2	2	2
3.	Amino acids Proteins	1	1	1
4.	Vitamins and Minerals	1	--	2
5.	Blood composition	1	2	1
6.	Practical	--	2	--

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Dept of Chemistry Inter Disciplinary Add On Course

Clinical Lab Technology-Syllabus

(w.e.f 2019-20 revised on October-2018)

Time: 3 Hours

Model Paper

Max. Mark: 75

SECTION – A

Answer any FIVE questions. Each question carries 2 Marks (5 X 2 = 10 Marks)

1. Define Molarity?
2. Write Molisch test?
3. What is a lipid?
4. What are amino acids?
5. Write the deficiency disease of vitamin A?
6. What is Plasma?

SECTION – B

Write any FIVE questions. Each question carries 5 Marks. (5 X 5 = 25 Marks)

7. Explain about Lipid profile test?
8. Explain about Liver function test?
9. Explain various concentrations methods of solutions?
10. Explain about Blood sugar test?
11. What is Hemoglobin? Explain its functions?
12. Write about Blood clotting and role of vitamin in it?
13. What are amino acids? Explain their classification?
14. Explain about serum creatine and serum urea test?

SECTION – C

Answer any FOUR questions. Each question carries 10 Marks. (4 X 10 = 40 Marks)

15. What are carbohydrates? Explain their classifications and biological importance?
16. Explain the functions, biological importance and normal levels of Cholesterol?
17. What are Proteins? Explain their functions and tests to identify?
18. Explain the sources and deficiency diseases of vitamin A, B, C, D, E.
19. Explain the sources and biological importance of Na, K and Ca?
20. Explain the composition of Blood and its function?